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DESCRIPTION

BISMUTH GLASS COMPOSITION, AND MAGNETIC HEAD AND PLASMA DISPLAY PANEL USING THE SAME AS SEALING MEMBER

Technical Field

The present invention relates to a low softening point bismuth glass composition mainly used for bonding, sealing and coating ceramics, glass and metal. The bismuth glass composition of the present invention is suitable for bonding a pair of magnetic core halves of a magnetic head or bonding a front plate and a rear plate of a plasma display panel.

Background Art

Ceramics, glass and metal, for example, have been used as materials for electronic equipment and various glass compositions have been used as materials for bonding, sealing or coating them. The glass compositions have various forms such as bulk, powder, fiber and a thin film. There are also a material consisting of the glass composition only, and a composite material comprising a combination of the glass composition and other materials. Further, to give various functions depending on its applications, the glass composition may be dispersed in a vehicle together with other materials and an appropriate filler so that it can be used as a paste.

The thus obtained paste may be used as a sealing member for magnetic heads, CRTs, liquid crystal display panels and plasma display panels.

Hereinafter, as a concrete example of applications of the sealing member comprising the glass composition, an explanation is given of the magnetic head and the plasma display panel.

First, a conventional magnetic head is explained.

A magnetic head is a device that performs recording and reproducing of magnetic information to and from a magnetic recording medium. The magnetic head comprises a pair of magnetic core halves, at least one of which being provided with a coil groove, and a magnetic gap member made of a non-magnetic material. The pair of magnetic core halves are abutted to each other via the magnetic gap member and bonded with a sealing member. As such a sealing member, a glass composition is used. The glass composition is an important constituent having an influence even on the properties of the magnetic head.

As a material for the magnetic core halves, ferrite has widely been used due to its excellent magnetic properties, wear resistance and machinability. A magnetic head manufactured with ferrite may be called a ferrite head.

In recent years, according to reduction in size and increase in capacity of the magnetic recording/reproducing device, a magnetic recording medium having high coercivity has

come to be used. Aiming at a magnetic head for high density magnetic recording, which is highly capable of writing signals to such a medium, an extension of the ferrite head has been under development.

For example, a magnetic head called an MIG (metal in gap) head is formed by applying magnetic metal films having high saturated magnetic flux density to the surfaces of the magnetic core halves to be opposed to a magnetic gap, abutting the surfaces facing to the magnetic gap to each other via the magnetic gap member, and bonding them with a sealing member. As the magnetic metal film, for example, a thin film of a magnetic metal material such as Fe-Ta-N, Fe-Nb-N, Fe-Nb-Si-B-N, Fe-Ta-C, Co-Ta-Zr-Nb or Co-Nb-Zr-N is used.

There is also a magnetic head comprising magnetic core halves, each of which comprising a magnetic metal film sandwiched between a pair of nonmagnetic substrates. The magnetic core halves are arranged such that the magnetic metal films are abutted at their end faces and bonded with a sealing member via a magnetic gap member. Such a magnetic head is called a stacked head.

In the process of manufacturing the magnetic head, the sealing member is used by softening, liquidizing, cooling and solidifying the glass composition by an appropriate heat treatment. Therefore, to prevent deterioration of the magnetic substance composing the magnetic head by heat, a glass composition usable at a temperature as low as possible

has been demanded. As such a glass composition that meets the demand, there is a so-called low softening point glass.

In general, the glass composition is apt to have a higher thermal expansion coefficient when the softening point is lower. However, in order to avoid breaking or cracking due to warpage after cooling, it is necessary to control the thermal expansion coefficient of the sealing member not to increase. Further, to exhibit optimum magnetic recording properties, it is necessary to control the warpage of the magnetic substance caused by a difference in thermal expansion coefficient from the sealing member. Therefore, there has been a demand for a sealing member having a thermal expansion coefficient suitable for the specifications of various magnetic heads.

The sealing members of the magnetic heads are required to have temperature properties and thermal expansion coefficients suitable for their applications, respectively. The sealing member used for the magnetic heads such as the ferrite head, MIG head and stacked head is required to have a working temperature of 450 to 650°C and a thermal expansion coefficient of 70×10^{-7} to $130 \times 10^{-7}/^{\circ}\text{C}$. For example, as to the sealing member for the MIG head, it is preferred that the working temperature is about 500°C and the thermal expansion coefficient is 75×10^{-7} to $100 \times 10^{-7}/^{\circ}\text{C}$.

The working temperature mentioned herein is a temperature at which the viscosity of the glass composition

becomes 10^3 Pa's. Further, the softening point is a temperature obtained by measurement based on a JIS test method R3103-1, which is determined as a temperature at which the viscosity of the glass composition becomes $10^{6.6}$ Pa's. With the exception of the case where the viscosity behavior of the glass composition is unique, a glass composition having a low softening point is apt to have a low working temperature.

Next, a common plasma display panel (hereinafter may be referred to as PDP) is explained.

FIG. 1 is an example of a partial schematic oblique view illustrating a major structure of a common PDP.

In FIG. 1, the PDP comprises a front plate 1 and a rear plate 8 bonded to each other.

The front plate 1 includes a front glass substrate 2, display electrodes 5 formed on a single surface thereof, each of which comprising a transparent conductive film 3 and a bus electrode 4, a dielectric layer 6 made of dielectric glass covering the display electrodes 5 and a dielectric layer protecting layer 7 made of magnesium oxide.

The rear plate 8 includes a rear glass substrate 9, address electrodes 10 formed on a single surface thereof, a dielectric layer 11 covering the address electrodes 10, barrier ribs 12 arranged at regular intervals on the top surface of the dielectric layer 11 to form discharge spaces 14 and fluorescent layers formed inside the discharge spaces 14 for color display. The fluorescent layers include a red

fluorescent layer 13(R), a green fluorescent layer 13(G) and a blue fluorescent layer 13(B) arranged sequentially, which are excited by ultraviolet light having a short wavelength of 147 nm, for example, to emit light.

The front plate 1 and the rear plate 8 are arranged such that the lengthwise directions of the display electrodes 5 and the address electrodes 10 are orthogonal to each other, and then bonded with a sealing member comprising a glass composition. FIG. 2 shows the arrangement of the display electrodes and the address electrodes, with an illustration of a sealing part.

Referring to FIG. 2, a sealing part 15 is the periphery of an overlapped region between the stacked front plate 1 and the rear plate 8. A paste comprising the sealing member is applied to the sealing part 15 and the front plate 1 and the rear plate 8 are bonded, thereby the inside thereof is sealed. Then, the display electrodes 5 and the address electrodes 10 are connected to external driving circuits (not shown), respectively.

FIG. 3 is an example of a cross section cut along the address electrode of the PDP shown in FIG. 1 illustrating the vicinity of the sealing part.

At least one of the front plate 1 and the rear plate 8 is provided with an air hole 17 for exhausting air from the inside of the PDP and introducing discharge gas therein. An end of a glass tube 18 communicating with the air hole 17 is

fixed to the periphery of the opening end of the air hole 17 with a sealing member 16. Then, while heating to a predetermined temperature, air is exhausted from the inside of the PDP. After the exhaustion is completed, the discharge gas is introduced up to a predetermined pressure. Then, finally, a basal portion of the glass tube 18 is heat-sealed.

In the thus manufactured PDP, discharge is performed for a predetermined time to stabilize light-emitting properties and discharge properties as required.

In the process of manufacturing the PDP, the heat treatment in the sealing step using the sealing member 16 needs to be conducted at low temperature so as not to cause any trouble to the other members. That is, since the PDP contains glass materials in the glass substrates, electrodes, dielectric layers and barrier ribs, it is necessary to prevent these glass materials from being softened or deformed by heating in the sealing step. Further, it is also necessary to prevent deterioration of the dielectric layer protecting layer 7 and the fluorescent layers 13(R), 13(G) and 13(B). In general, it is desirable to carry out the sealing step at a treatment temperature of 500°C or lower.

Thus, in the step of sealing the magnetic head or the PDP, it is preferable to use the sealing member at low temperature. Therefore, low softening point glass is used for the glass composition.

For example, examples of the low softening point

glass used for the sealing member for the magnetic head include lead glasses such as $\text{SiO}_2\text{-B}_2\text{O}_3\text{-PbO}$ glass (Japanese Laid-Open Patent Publication No. HEI 8-180310) and $\text{B}_2\text{O}_3\text{-PbO-ZnO}$ glass. Further, as the sealing glass for the PDP, lead glass containing PbO as a main component is used. That is, it is essential to contain lead to achieve the low softening point.

Although lead is heavily used as a component of the low softening point glass, it has been pointed out that lead has toxicity to human beings and hazardous properties to environment. Further, working environment in the manufacture of the magnetic head or the PDP and influence on the environment that accompanies disposal of the products have been brought into question. Therefore, there has been a demand for a sealing member using a lead free glass composition.

As the lead free low-melting glass, phosphate glass has been developed. However, in actual use, it is not sufficiently reliable, particularly with respect to water resistance. Moreover, a sealing member comprising a glass composition having low water resistance, for example, is apt to absorb moisture in the atmosphere in the manufacturing process of the PDP and the moisture may possibly remain in the PDP to cause harmful effect on display performance. In order to prevent this, the glass composition contained in the sealing member is required to be excellent in water resistance.

Further, there has also been studied a glass

composition comprising lead free, low-melting point bismuth glass (Japanese Laid-Open Patent Publication No. HEI 10-139478).

Disclosure of Invention

In view of the above, the present invention provides a lead free, highly reliable bismuth glass composition having a low working temperature, an appropriate thermal expansion coefficient and mechanical strength, as well as excellent water resistance. Further, an object of the present invention is to provide a magnetic head and a PDP using a sealing member comprising the bismuth glass composition that reduces load to the environment.

More specifically, the present invention provides a bismuth glass composition comprising 0.5 to 14 wt% of SiO_2 , 3 to 15 wt% of B_2O_3 , 4 to 22 wt% of ZnO , 55 to 90 wt% of Bi_2O_3 and 4 wt% or less of Al_2O_3 , and further comprising 5 wt% or less of an oxide of Group A, 12 wt% or less of an oxide of Group B and 0.1 to 10 wt% of an oxide of Group C, wherein the oxide of Group A is at least one selected from the group consisting of Li_2O , Na_2O and K_2O , the oxide of Group B is at least one selected from the group consisting of MgO , CaO , SrO and BaO , and the oxide of Group C is at least one selected from the group consisting of Sc_2O_3 , Y_2O_3 , La_2O_3 , CeO_2 , Pr_2O_3 , Nd_2O_3 , Sm_2O_3 , Eu_2O_3 , Gd_2O_3 , Tb_2O_3 , Dy_2O_3 , Ho_2O_3 , Er_2O_3 , Tm_2O_3 , Yb_2O_3 and Lu_2O_3 .

The glass composition preferably comprises 0.5 to 12 wt% of SiO_2 , 3 to 9 wt% of B_2O_3 , 4 to 19 wt% of ZnO , 55 to 85 wt% of Bi_2O_3 and 0.1 to 4 wt% of Al_2O_3 .

In the glass composition, the content of the oxide of Group A is preferably 4 wt% or less. Further, the oxide of Group A preferably comprises 2 wt% or less of Li_2O , 3 wt% or less of Na_2O and 4 wt% or less of K_2O .

In the glass composition, the content of the oxide of Group B is preferably 10 wt% or less. Further, the oxide of Group B preferably comprises 5 wt% or less of MgO , 6 wt% or less of CaO , 8 wt% or less of SrO and 10 wt% or less of BaO .

In the glass composition, the content of the oxide of Group C is preferably 0.1 to 5 wt%.

The weight ratio of ZnO to B_2O_3 is preferably 0.8 to 2.8. In particular, if the glass composition is used as the sealing member for the PDP, the weight ratio is preferably 1.1 to 2.5.

The weight ratio of Al_2O_3 to SiO_2 is preferably 0.5 or less.

If such a glass composition is used as the sealing member for the magnetic head, the glass composition preferably comprises 1.7 to 12 wt% of SiO_2 , 3 to 9 wt% of B_2O_3 , 9.5 to 19 wt% of ZnO , 62 to 80 wt% of Bi_2O_3 and 0.1 to 4 wt% of Al_2O_3 .

Further, the present invention relates to a magnetic head comprising a pair of magnetic core halves, at least one of which being provided with a coil groove, a magnetic gap

member interposed between surfaces of the magnetic core halves facing to a magnetic gap, and the sealing member for the magnetic head for bonding the pair of magnetic core halves. A magnetic metal film may be formed on at least one of the surfaces facing to the magnetic gap. Alternatively, each of the pair of magnetic core halves comprises a pair of nonmagnetic substrates and a magnetic metal film sandwiched between the nonmagnetic substrates.

Still further, the present invention relates to a magnetic recording/reproducing device comprising the magnetic head to perform recording and reproducing information to and from a magnetic information recording medium.

If the above-described glass composition is used as the sealing member for the PDP, the glass composition preferably comprises 1.1 to 4.5 wt% of SiO_2 , 4 to 9 wt% of B_2O_3 , 9.5 to 18 wt% of ZnO and 72 to 85 wt% of Bi_2O_3 . Further, the glass composition used as the sealing member for the PDP preferably comprises 0.1 to 2 wt% of Al_2O_3 . Still further, the glass composition used as the sealing member for the PDP preferably comprises 8 wt% or less of the oxide of Group B. In particular, the oxide of Group B preferably comprises 2 wt% or less of MgO , 0.1 to 4.5 wt% of CaO , 0.1 to 4.5 wt% of SrO and 4 wt% or less of BaO .

Preferably, the sealing member for the PDP further comprises a low-expansion ceramic filler in a weight ratio of 0.01 to 4 with respect to the bismuth glass composition. As

the low-expansion ceramic filler, it is preferable to use at least one selected from the group consisting of cordierite, willemite, forsterite, anorthite, zircon, mullite, β -eucryptite, β -spodumene, cristobalite, barium titanate, titanium oxide, tin oxide, aluminum oxide, zirconium oxide and quartz glass.

Moreover, the present invention relates to a PDP comprising a front plate and a rear plate facing to each other, display electrodes and address electrodes arranged between the front plate and the rear plate, barrier ribs for isolating the address electrodes, dielectric layers covering the surfaces of the display electrodes and the address electrodes, respectively, and said sealing member for the PDP for bonding the peripheries of the front plate and the rear plate. The present invention also relates to a PDP provided with an air hole formed in the front plate or the rear plate, further comprising a glass tube communicating with the air hole, wherein an opening end of the air hole and an end of the glass tube are bonded with the sealing member for the PDP.

In the present invention, the content of each element in the glass composition is expressed by the ratio of a certain oxide. However, it is not necessary that the elements exist in the form of oxide. The content of the oxide shown in the present invention is based on the assumption that every element contained in the glass composition forms the oxide as indicated.

Brief Description of Drawings

FIG. 1 is an example of a partial oblique sectional view illustrating a major structure of a common PDP.

FIG. 2 is a view illustrating the arrangement of display electrodes and address electrodes of the PDP and a sealing part.

FIG. 3 is an example of a section cut along the address electrode of the PDP shown in FIG. 1 illustrating the vicinity of the sealing part.

FIG. 4 is a schematic oblique view illustrating an example of a ferrite head according to an embodiment of the present invention.

FIG. 5 is a schematic oblique view illustrating an example of an MIG head according to an embodiment of the present invention.

FIG. 6 is a schematic oblique view illustrating an example of a stacked head according to an embodiment of the present invention.

FIG. 7 is a view illustrating an example of a surface of the stacked head of FIG. 6 that slides on a magnetic recording medium.

FIG. 8 is an oblique view illustrating an example of a rotating drum of a magnetic recording/reproducing device according to an embodiment of the present invention.

FIG. 9 is a schematic view illustrating an example

of an inner mechanism of the magnetic recording/reproducing device having the rotating drum of FIG. 8.

Best Mode for Carrying Out the Invention

The bismuth glass composition of the present invention comprises 0.5 to 14 wt% of SiO_2 , 3 to 15 wt% of B_2O_3 , 4 to 22 wt% of ZnO , 55 to 90 wt% of Bi_2O_3 and 4 wt% or less of Al_2O_3 . The Al_2O_3 content may be 0 %.

If the content of SiO_2 in the glass composition is less than 0.5 wt%, the glass composition cannot be obtained with stability. If the content exceeds 14 wt%, the working temperature becomes high. The SiO_2 content is preferably 0.5 to 12 wt%. Further, if the content of B_2O_3 is less than 3 wt%, the thermal expansion coefficient becomes high. If the content exceeds 15 wt%, the glass composition cannot be obtained with stability. The B_2O_3 content is preferably 3 to 9 wt%. Further, if the content of ZnO is less than 4 wt%, water resistance is lowered. If the content exceeds 22 wt%, the glass composition cannot be obtained with stability. The ZnO content is preferably 4 to 19 wt%. Further, if the content of Bi_2O_3 is less than 55 wt%, the working temperature becomes high. If the content exceeds 90 wt%, the glass composition cannot be obtained with stability. The Bi_2O_3 content is preferably 55 to 85 wt%. Moreover, if the content of Al_2O_3 exceeds 4 wt%, the working temperature becomes high. The Al_2O_3 content is preferably 0.1 to 4 wt%.

The glass composition further comprises 5 wt% or less of an oxide of Group A, 12 wt% or less of an oxide of Group B and 0.1 to 10 wt% of an oxide of Group C. The content of the oxide of Group A may be 0 %. Further, the content of the oxide of group B may be 0 %. The oxide of Group A is at least one selected from the group consisting of Li_2O , Na_2O and K_2O . The oxide of Group B is at least one selected from the group consisting of MgO , CaO , SrO and BaO . Further, the oxide of Group C is at least one selected from the group consisting of Sc_2O_3 , Y_2O_3 , La_2O_3 , CeO_2 , Pr_2O_3 , Nd_2O_3 , Sm_2O_3 , Eu_2O_3 , Gd_2O_3 , Tb_2O_3 , Dy_2O_3 , Ho_2O_3 , Er_2O_3 , Tm_2O_3 , Yb_2O_3 and Lu_2O_3 .

Among the oxides of Group C, in particular, Nd_2O_3 , Gd_2O_3 and Er_2O_3 are preferably used. The content of the oxide of Group C is preferably 0.1 to 5 wt%.

Since the glass composition does not contain lead, pollution by lead in the working environment decreases and the lead content in waste generated in the manufacture process is reduced.

Al_2O_3 functions to improve the water resistance and stabilizes the glass composition in the manufacture thereof, thereby controlling crystal precipitation in the glass in heat treatment after the manufacture of the glass composition. The oxide of Group A lowers the softening point of the glass composition, thereby decreasing the working temperature of the glass composition. The oxide of Group B stabilizes the glass composition. The oxide of Group C improves mechanical

strength of the glass composition.

In the glass composition according to the present invention, the content of the oxide of Group A is preferably 4 wt% or less. In particular, it is preferred that the oxide of Group A comprises 2 wt% or less of Li_2O , 3 wt% or less of Na_2O and 4 wt% or less of K_2O . In this case, the content of at least one or two selected from the group consisting of Li_2O , Na_2O and K_2O may be 0 %. Among them, if there is a particular demand for maintaining the glass composition with stability, Na_2O is preferably used. If the Li_2O content exceeds 2 wt%, the Na_2O content exceeds 3 wt% or the K_2O content exceeds 4 wt%, the glass composition tends to fail to maintain its hermeticity.

In the glass composition according to the present invention, the content of the oxide of Group B is preferably 10 wt% or less. In particular, the oxide of Group B preferably comprises 5 wt% or less of MgO , 6 wt% or less of CaO , 8 wt% or less of SrO and 10 wt% or less of BaO . In this case, the content of at least one, two or three selected from the group consisting of MgO , CaO , SrO and BaO may be 0 %. Among them, if there is a particular demand for stabilizing the glass composition, CaO or SrO is preferably used. If the MgO content exceeds 5 wt%, the CaO content exceeds 6 wt%, the SrO content exceeds 8 wt% or the BaO content exceeds 10 wt%, the glass composition tends to fail to maintain its hermeticity.

A glass composition comprising 1.7 to 12 wt% of SiO_2 , 3 to 9 wt% of B_2O_3 , 9.5 to 19 wt% of ZnO , 62 to 80 wt% of Bi_2O_3 and 0.1 to 4 wt% of Al_2O_3 , which is an embodiment of the above-described glass composition, is suitable for a sealing member for a magnetic head.

Further, as another embodiment of the glass composition, a glass composition comprising 1.1 to 4.5 wt% of SiO_2 , 4 to 9 wt% of B_2O_3 , 9.5 to 18 wt% of ZnO and 72 to 85 wt% of Bi_2O_3 is suitable for a sealing member for a PDP. The glass composition used as the sealing member of the PDP preferably comprises 0.1 to 2 wt% of Al_2O_3 . Further, the content of the oxide of Group B is preferably 8 wt% or less. In particular, the oxide of Group B preferably comprises 2 wt% or less of MgO , 0.1 to 4.5 wt% of CaO , 0.1 to 4.5 wt% of SrO and 4 wt% or less of BaO . In this case, the content of one or two selected from the group consisting of MgO and BaO may be 0 %. If the MgO content exceeds 2 wt%, the glass composition tends to fail to maintain its hermeticity. If the CaO content exceeds 4.5 wt% and the SrO content exceeds 4.5 wt%, the water resistance is apt to decrease. Further, if the BaO content exceeds 4 wt%, the glass composition tends to fail to maintain its hermeticity.

In the glass composition according to the present invention, the weight ratio of ZnO to B_2O_3 is preferably 0.8 to 2.8. In particular, if it is used as the sealing member for the PDP, the weight ratio is preferably 1.1 to 2.5. If the

weight ratio of ZnO to B₂O₃ decreases, the hermeticity of the glass composition cannot be maintained. Further, if the weight ratio increases, the glass composition tends to fail to maintain the hermeticity.

In the glass composition according to the present invention, the weight ratio of Al₂O₃ to SiO₂ is preferably 0.5 or less. If the weight ratio of Al₂O₃ to SiO₂ increases, the glass composition tends to fail to maintain its hermeticity.

Meeting these conditions allows stabilization of the glass composition in the manufacture thereof. Further, even when the glass composition is subjected to heat treatment after the manufacture thereof, crystal precipitation in the glass is controlled.

The glass composition according to the present invention may be used for bonding, sealing or covering ceramics, glass and metal. Further, it may also be used as a paste material having various functions. For example, the glass composition of the present invention may be used in place of glass materials that have conventionally been used in every application, such as various parts of electronic equipment. More specifically, it may be used for various LCR parts, semiconductor packages, other electronic parts, and display devices such as CRTs, liquid crystal display panels and plasma display panels. Further, it may also be used for tube products for lighting, enameled products and pottery products.

Hereinafter, with reference to the drawings, an explanation is given of a magnetic head, a magnetic recording/reproducing device and a PDP using the sealing member comprising the bismuth glass composition of the present invention.

(1) Magnetic head

FIG. 4 shows a schematic oblique view illustrating an example of a ferrite head according to an embodiment of the present invention. A ferrite head 20 comprises a pair of magnetic core halves 22 and 23 made of ferrite, at least one of which being provided with a coil groove 21, sealing members 25 and 26 for bonding the magnetic core halves 22 and 23, and a magnetic gap member 24 interposed between the surfaces of the magnetic core halves 22 and 23 facing to a magnetic gap. The surfaces of the paired magnetic core halves 22 and 23 facing to the magnetic gap are abutted to each other via the magnetic gap member 24 and the magnetic gap is formed on the interface between the abutted surfaces. The sealing members 25 and 26 bond and unite the magnetic core halves 22 and 23.

As a magnetic head according to another embodiment of the present invention, FIG. 5 shows a schematic oblique view of an example of an MIG head. An MIG head 30 comprises a pair of magnetic core halves 32 and 33 made of ferrite, at least one of which being provided with a coil groove 31, magnetic metal films 37 and 38 formed on the surfaces of the magnetic core halves 32 and 33 facing to a magnetic gap,

sealing members 35 and 36 for bonding the magnetic core halves 32 and 33 and a magnetic gap member 34 interposed between the magnetic metal films 37 and 38. The surfaces of the pair of magnetic core halves 32 and 33 facing to the magnetic gap are abutted to each other via the magnetic gap member 34 and the magnetic gap is formed on the interface between the abutted surfaces. The sealing members 35 and 36 bond and unite the magnetic core halves 32 and 33.

As a magnetic head according to still another embodiment of the present invention, FIG. 6 shows a schematic oblique view illustrating an example of a stacked head. A stacked head 40 comprises a pair of multilayer magnetic core halves 42 and 43, at least one of which being provided with a coil groove 41, sealing members 45 to 47 for bonding the multilayer magnetic core halves 42 and 43 and a magnetic gap member 44 interposed between the surfaces of the multilayer magnetic core halves 42 and 43 facing to a magnetic gap. FIG. 7 shows a surface of the stacked head 40 that slides on a magnetic recording medium. The multilayer magnetic core halves 42 and 43 comprises magnetic metal films 53 and insulating films 54 stacked alternately in a thickness direction thereof and a pair of nonmagnetic substrates 51 and 52 sandwiching the films.

The sealing members used in the above-described magnetic heads are made of a bismuth glass composition, which is molded into a circular rod or a square rod for use. In the

step of sealing the magnetic head, first, a coil groove or a hollow groove for embedding the sealing member therein is formed in the pair of magnetic core halves, and then a film of the magnetic gap member is formed on the surfaces facing to the magnetic gap. Then, they are abutted to each other and the sealing member is arranged in the hollow groove.

Subsequently, the sealing member is softened and liquidized by an appropriate heat treatment and then cooled and solidified to bond the pair of core halves. The thus obtained magnetic core block is cut into a predetermined size and polished to form a magnetic head chip, to which a base is bonded and a coil is provided, thereby completing the magnetic head.

It is necessary to control the temperature of the heat treatment in the sealing step not to exceed the resistible temperature of the material used for each magnetic head. The heat treatment is preferably performed at 600 to 700°C to form the ferrite head, 500 to 600°C to form the MIG head and 450 to 500°C to form the stacked head.

Regarding the above-described sealing step, a glass composition which can be used at a temperature as low as possible has been demanded so as not to deteriorate the magnetic substance composing the magnetic head by heat. For this reason, the sealing member used in a common magnetic head preferably has a working temperature of 450 to 650°C and a thermal expansion coefficient of 70×10^{-7} to $130 \times 10^{-7}/^{\circ}\text{C}$.

As to the magnetic cores, magnetic metal films,

magnetic gap members, insulating films and nonmagnetic substrates of the above-described ferrite head 20, MIG head 30 and stacked head 40, conventionally used materials may be used without any limitation.

The sealing member for the magnetic head according to the present invention may also be used for magnetic heads having other structures than those of the above-described magnetic heads.

(2) Magnetic recording/reproducing device

Next, an explanation is given of an embodiment of a magnetic recording/reproducing device according to the present invention. FIG. 8 is an oblique view illustrating an example of a rotating drum of a magnetic recording/reproducing device. FIG. 9 is a schematic view illustrating an inner mechanism of an example of a magnetic recording/reproducing device having the rotating drum of FIG. 8.

A rotating drum 61 shown in FIG. 8 comprises a lower drum 62, an upper rotating drum 63, and a magnetic head 64 provided on the periphery thereof. A magnetic tape, which is not shown, travels along a lead 65 with inclination relative to the rotation axis of the upper rotating drum 63. The magnetic head 64 slides with inclination relative to the traveling direction of the magnetic tape. Further, a plurality of grooves 66 are formed on the periphery surface of the upper rotating drum 63 such that the upper rotating drum 63 and the magnetic tape can slide in close contact with

stability. The grooves 66 emit air caught up between the magnetic tape and the upper rotating drum 63.

As shown in FIG. 9, a magnetic recording/reproducing device 100 comprises a rotating drum 61, a supply reel 101, a take-up reel 102, rotating posts 103, 104, 105, 106, 107 and 108, inclined posts 109 and 110, a capstan 111, a pinch roller 112 and a tension arm 113. The magnetic heads 64 for recording/reproducing are arranged on the periphery surface of the rotating drum 61.

A magnetic tape 114 wound on the supply reel 101 travels between the pinch roller 112 and the capstan 111 to be wound on the take-up reel 102. The rotating drum 61 is of an upper rotating drum type, in which two magnetic heads 64 are arranged such that they are protruded by about 20 μm from the periphery surface of the rotating drum.

According to the present embodiment, the rotating drum of the upper rotating drum type is exemplified. However, the present invention is also applicable to a magnetic recording/reproducing device of a center drum rotating type adopting three drums of an upper drum, a center drum and a lower drum. Further, not only the magnetic head for recording/reproducing, the present invention is also applicable to a magnetic head for recording and a magnetic head for reproducing. In the recording/reproducing device, the number of the magnetic head is not limited. Further, the explanation herein is directed to the case where the magnetic

tape is a recording medium. However, the present invention is also applicable to a magnetic head which performs recording/reproducing information to and from a disc-shaped recording medium, as well as a magnetic recording/reproducing device comprising the same.

(3) PDP

Hereinafter, an explanation is given of a PDP using the bismuth glass composition of the present invention as the sealing member with reference to the drawings. The PDP described herein is the same in structure as that shown in Figs. 1 to 3 described in Background Art except that the composition of the glass composition in the sealing member is varied.

First, an example of a method for forming a front plate 1 of the PDP is explained with reference to FIG. 1.

First, onto a front glass substrate 2 made of high distortion point glass, a paste for forming transparent conductive films is applied by screen printing. Then, a paste for forming bus electrodes as a supplement to the conductivity is applied thereto by screen printing. These pastes are baked to form display electrodes 5. The transparent conductive film 3 is made of ITO (indium tin oxide), for example, and the bus electrodes 4 are made of, for example, an Ag film or a three-layered film of Cr/Cu/Cr. Alternative to the screen printing, the transparent conductive films 3 and the bus electrodes 4 may be formed by photo-etching.

A paste containing dielectric glass is applied to the top of the display electrodes 5 by screen printing and baked to form a dielectric layer 6. The dielectric layer 6 may be formed by disposing a sheet containing dielectric glass and a resin on the display electrodes 5, followed by baking. The sheet as described above may be obtained by mixing the dielectric glass and the resin with a solvent, forming a coating film of the mixture and volatilizing the solvent from the film.

On the surface of the dielectric layer 6, a dielectric layer protecting layer 7 made of magnesium oxide, for example, is formed by vapor deposition or CVD. In this manner, the front plate 1 is formed.

Next, a method for forming a rear plate 8 of the PDP is explained with reference to FIG. 1.

First, on a rear glass substrate 9 made of high distortion point glass, a paste for forming electrodes (e.g., Ag or Cr/Cu/Cr) is applied in a stripe pattern at regular intervals by screen printing or the like and baked to form address electrodes 10.

Then, a dielectric layer 11 is formed to cover the top of the address electrodes 10. The dielectric layer 11 is formed by applying a paste containing dielectric glass to the address electrodes 10 by screen printing, followed by baking.

On the address electrodes 10, barrier ribs 12 serving as partitions among discharge spaces 14 are provided

in a stripe pattern. The barrier ribs 12 are formed by, for example, applying a paste containing a glass material repeatedly by screen printing, followed by baking. Alternatively, the address electrodes 10, the dielectric layer 11 and the barrier ribs 12 may be formed by photo-etching or the like.

Inside the thus formed discharge spaces 14, pastes each containing fluorescent materials of red (R), green (G) and blue (B) are applied sequentially by screen printing. By baking them, a red fluorescent layer 13(R), a green fluorescent layer 13(G) and a blue fluorescent layer 13(B) are formed. As examples of the fluorescent materials used, $Y_2O_3:Eu$ is used for the red fluorescent layer 13(R), $Zn_2SiO_4:Mn$ is used for the green fluorescent layer 13(G) and $BaMgAl_{10}O_{17}:Eu$ is used for the blue fluorescent layer 13(B).

With reference to Figs. 2 and 3, an explanation is given of bonding of the thus formed front plate 1 and rear plate 8 using a sealing member.

A sealing member 16 comprises bismuth glass composition powder, which is kneaded with a resin and a solvent to prepare a paste. Examples of the resin include ethyl cellulose, nitrocellulose, acrylic resin and the like. Examples of the solvent include terpeneol, isoamyl acetate and the like. The amounts of the resin and the solvent contained in the paste are preferably 0.01 to 30 parts by weight and 3 to 120 parts by weight, respectively, relative to 100 parts by

weight of the glass composition.

The paste is applied to a sealing part 15 of the rear plate 8 shown in FIG. 2 by screen printing or injection. Then, the whole is provisionally baked at a temperature that allows removal of the resin component contained in the paste (350°C for example). Then, the front plate 1 and the rear plate 8 are stacked such that the display electrodes 5 and the address electrodes 10 are orthogonal to each other. Then, the sealing member 16 is further baked. In this manner, bonding of the front plate 1 and the rear plate 8 is achieved.

The sealing member 16 is applied to at least either one of the front plate 1 and the rear plate 8 and the sealing method is not limited to the above.

Next, with reference to FIG. 3, filling of gas into the discharge spaces 14 is explained.

First, an opening end of an air hole 17 formed in the front plate 1 or the rear plate 8 is bonded with an end of a glass tube 18 communicating with the air hole 17. For example, a paste comprising the sealing member is applied to the opening end of the air hole 17 and the end of the glass tube 18 by injection, followed by baking, to achieve the bonding.

Subsequently, in order to vacuum the discharge spaces 14, air is exhausted through the glass tube 18 while heating the whole of the bonded front plate 1 and rear plate 8, thereby purging gas existing between the front plate and the

rear plate. Then, discharge gas is filled therein through the glass tube 18 such that the pressure in the discharge spaces 14 reaches a predetermined level. Then, the glass tube 18 is heated and fused to cut it at a basal portion thereof, thereby sealing the air hole 17. Finally, the display electrodes 5 and the address electrodes 10 are connected to external driving circuits (not shown) to complete the PDP.

The bonding of the air hole 17 and the glass tube 18 may be performed at the same time as the bonding of the front plate 1 and the rear plate 8.

In the above-described sealing step, it is necessary not to cause harmful effects on the components of the PDP by heat such as deterioration and degradation. For that purpose, in a common PDP, the sealing is preferably conducted at a temperature not higher than 500°C.

The sealing with the sealing member 16 is not limited to the above-described method of using the paste prepared by kneading the bismuth glass composition powder, the resin and the solvent. For example, the sealing member 16 may be formed into a predetermined shape such as a rod or a frame, which is disposed on the sealing part, followed by heating and softening, to perform the same sealing step as described above.

In the present embodiment, the explanation is given by way of a structure of a PDP of alternating current system. However, the discharge system of the PDP to which the present invention is applicable is not limited thereto. The present

invention can be applied to both of the alternate current system and the direct current system.

Hereinafter, the present invention is explained in further detail by way of examples, but the present invention is not limited to these examples.

Example 1

Bismuth glass compositions having the compositions shown in Tables 1 to 20 were formed.

After the predetermined materials were prepared and mixed, the obtained mixture was placed in a platinum crucible and melted in an electric furnace at 900 to 1300°C for 1 hour. Then, the obtained molten glass was quenched with a roller to form a flaky glass composition. Tables 1 to 20 show the composition, glass stability, working temperature, thermal expansion coefficient, strength and water resistance of the obtained glass compositions.

The glass stability, working temperature, thermal expansion coefficient, strength and water resistance were evaluated in the following manner.

[Stability]

The glass stability was judged from whether or not precipitation of crystals (hereinafter may be referred to as devitrification) was caused from the glassy state in heat treatment performed at each working temperature for 30 minutes. The stability was indicated by ○ when the glass composition

was not devitrified in the manufacture thereof and did not cause the crystal precipitation in the heat treatment at the above-described test temperature, Δ when the glass composition was not devitrified in the manufacture thereof but caused the crystal precipitation in the heat treatment at the above-described test temperature, or \times when the glass composition was devitrified in the manufacture thereof and the glass composition was not obtained. If the glass composition is used as a sealing member for a magnetic head, the glass composition is preferably amorphous in view of flowability. However, as long as there are no problems such as generation of cracks or decrease in strength, crystals may be precipitated in the sealing member after the sealing.

[Working temperature]

The working temperature of the glass composition was determined as a temperature at which the viscosity of the molten glass composition becomes 10^3 Pa's. If the glass composition is used as the sealing member for the magnetic head, the working temperature is preferably 450 to 650°C. Preferably, the sealing member is softened and liquidized at a temperature around the working temperature to seal the magnetic head. However, the sealing member may be used with a viscosity suitable for the manufacture method of the magnetic head.

[Thermal expansion coefficient]

The glass composition was re-fused, poured into a

mold and cooled to form a glass rod of 4 mm in diameter and 20 mm in length. Under heating at 10°C/min, linear expansivity of the glass rod was measured. Then, an average thermal expansion coefficient at 30 to 300°C was calculated. If the glass composition is used as the sealing member for the magnetic head, the average thermal expansion coefficient is preferably 70×10^{-7} to $130 \times 10^{-7}/^{\circ}\text{C}$.

[Mechanical strength]

The mechanical strength was measured by a three-point bending method. The glass composition was re-fused to be a melt having a suitable viscosity, from which a glass rod of 1 mm in diameter and 30 mm in length was drawn. The glass rod was held horizontally by two points at an interval of 20 mm at the center portion thereof, and a load was applied to the center between the two points from above by a load cell at a rate of 1 mm/min. Based on the load that broke the glass rod, the mechanical strength was calculated.

[Water resistance]

To evaluate the water resistance, a cube of 10 mm on a side was manufactured as a sample by re-fusing the glass composition, which was immersed into boiling ion-exchanged water for an hour. A value of decrease in weight per unit surface area of the glass composition was indicated as the water resistance by the unit of (mg/cm^2). If the glass composition is used as the sealing member for the magnetic head, this value is preferably 1.0 mg/cm^2 or

less.

Table 1 shows the glass compositions manufactured to make a study of the content of SiO_2 . The glass compositions of Nos. 2 to 6 are Examples and those of Nos. 1 and 7 are Comparative Examples.

Table 1

		Comparative Example	Example					Comparative Example
No.		1	2	3	4	5	6	7
Composition	SiO_2	0.3	0.5	1.7	3.2	12.0	14.0	15.0
(wt%)	B_2O_3	5.1	5.1	6.8	6.6	8.5	6.5	5.5
	ZnO	5.5	5.5	9.9	12.4	14.5	14.5	14.5
	Bi_2O_3	87.6	87.6	80.0	75.6	63.2	63.2	63.2
	Al_2O_3	0.1	—	0.4	1.0	0.1	0.1	0.1
	Na_2O	1.2	1.2	0.9	0.6	0.7	0.7	0.7
	La_2O_3	0.2	0.1	—	0.6	—	—	—
	CeO_2	—	—	—	—	1.0	1.0	1.0
	Er_2O_3	—	—	0.3	—	—	—	—
Stability		×	○	○	○	○	△	△
Working temperature (°C)		—	455	500	535	630	650	660
Thermal expansion coefficient ($\times 10^{-7}/^\circ\text{C}$)		—	130	110	98	78	81	80
Mechanical Strength (MPa)		—	116	121	128	139	144	151
Water resistance (mg/cm^2)		—	0.8	0.6	0.5	0.3	0.0	0.0

As clearly shown in Table 1, when the SiO_2 content was 0.3 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was not obtained with stability. When the SiO_2 content was 15 wt%, the working temperature exceeded 650°C . From the above results, the SiO_2 content is preferably 0.5 to 14 wt%.

Further, when the SiO_2 content is 0.5 wt%, the water resistance is close to 1 mg/cm^2 . When the SiO_2 content is 14 wt%, crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may cause a problem in maintaining hermeticity. Therefore, if the glass composition is used particularly as the sealing member for the magnetic head, the SiO_2 content is preferably 1.7 to 12 wt%.

Table 2 shows the glass compositions manufactured to make a study of the content of B_2O_3 . The glass compositions of Nos. 9 to 12 are Examples and those of Nos. 8 and 13 are Comparative Examples.

Table 2

		Comparative Example	Example				Comparative Example
No.		8	9	10	11	12	13
Composition	SiO ₂	2.5	2.0	4.3	10.0	5.8	5.8
(wt%)	B ₂ O ₃	2.5	3.0	7.5	9.0	15.0	16.4
	ZnO	7.0	8.0	12.8	20.3	13.4	13.4
	Bi ₂ O ₃	82.5	82.5	72.4	57.9	63.1	63.1
	Al ₂ O ₃	1.0	1.0	1.0	1.0	—	—
	Na ₂ O	1.0	1.0	—	—	—	1.0
	K ₂ O	—	—	—	1.0	—	—
	CaO	1.0	1.0	—	0.5	—	—
	SrO	1.0	1.0	—	—	—	—
	Y ₂ O ₃	—	—	—	—	0.2	—
	La ₂ O ₃	1.0	—	1.0	0.3	0.5	—
	CeO ₂	—	—	—	—	0.5	—
	Nd ₂ O ₃	—	—	—	—	0.5	0.3
	Gd ₂ O ₃	0.5	0.5	1.0	—	0.5	—
	Er ₂ O ₃	—	—	—	—	0.5	—
Stability		○	○	○	○	△	×
Working temperature(°C)		480	470	560	625	610	—
Thermal expansion coefficient (×10 ⁻⁷ /°C)		133	129	89	80	70	—
Mechanical Strength (MPa)		124	120	127	131	122	—
Water resistance (mg/cm ²)		0.5	0.7	0.5	0.2	0.0	—

As clearly shown in Table 2, when the B₂O₃ content was 2.5 wt%, the thermal expansion coefficient was $133 \times 10^{-7}/^{\circ}\text{C}$, which exceeded $130 \times 10^{-7}/^{\circ}\text{C}$. When the B₂O₃ content was 16.4 wt%, the glass composition was devitrified in the manufacture thereof and the glass composition was not obtained with stability. Therefore, the B₂O₃ content is preferably 3 to 15 wt%.

If the B_2O_3 content is 15.0 wt%, crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity. Therefore, if the glass composition is used particularly as the sealing member for the magnetic head, the B_2O_3 content is preferably 3 to 9 wt%.

Table 3 shows the glass compositions manufactured to make a study of the content of ZnO. The glass compositions of Nos. 15 to 19 are Examples and those of Nos. 14 and 20 are Comparative Examples.

Table 3

		Comparative Example	Example					Comparative Example
No.		14	15	16	17	18	19	20
Composition	SiO ₂	5.1	5.0	6.0	7.6	3.5	9.0	8.0
(wt%)	B ₂ O ₃	4.0	5.0	6.3	7.8	7.0	8.6	8.5
	ZnO	3.2	4.0	9.5	15.8	19.0	22.0	23.2
	Bi ₂ O ₃	82.5	80.9	73.0	62.3	69.5	57.9	56.5
	Al ₂ O ₃	1.4	1.5	1.3	1.0	0.4	1.0	1.8
	Na ₂ O	2.2	2.0	2.0	1.0	—	1.0	1.5
	K ₂ O	—	—	—	3.0	—	—	—
	MgO	—	0.5	—	—	—	—	—
	CaO	1.2	0.5	1.2	0.5	—	—	—
	Sc ₂ O ₃	—	—	—	—	—	0.5	—
	La ₂ O ₃	0.4	—	—	0.5	—	—	0.5
	CeO ₂	—	—	—	0.5	—	—	—
	Pr ₂ O ₃	—	0.1	—	—	—	—	—
	Nd ₂ O ₃	—	0.1	0.7	—	—	—	—
	Sm ₂ O ₃	—	0.1	—	—	—	—	—
	Eu ₂ O ₃	—	0.1	—	—	—	—	—
	Gd ₂ O ₃	—	0.1	—	—	—	—	—
	Tb ₂ O ₃	—	0.1	—	—	—	—	—
	Dy ₂ O ₃	—	—	—	—	0.1	—	—
	Ho ₂ O ₃	—	—	—	—	0.1	—	—
	Er ₂ O ₃	—	—	—	—	0.1	—	—
	Tm ₂ O ₃	—	—	—	—	0.1	—	—
	Yb ₂ O ₃	—	—	—	—	0.1	—	—
	Lu ₂ O ₃	—	—	—	—	0.1	—	—
Stability		○	○	○	○	○	△	×
Working temperature (°C)		520	525	565	610	540	615	—
Thermal expansion coefficient (×10 ⁻⁷ /°C)		118	116	109	101	82	80	—
Mechanical strength (MPa)		121	124	128	132	121	130	—
Water resistance (mg/cm ²)		1.5	0.5	0.3	0.2	0.2	0.1	—

As clearly shown in Table 3, when the ZnO content was 3.2 wt%, the water resistance was 1.5 mg/cm^2 , which exceeded 1.0 mg/cm^2 . When the ZnO content was 23.2 wt%, the glass composition was devitrified in the manufacture thereof, therefore the glass composition was not obtained with stability. Accordingly, the ZnO content is preferably 4 to 22 wt%.

If the glass composition is used particularly as the sealing member for the magnetic head, the ZnO content is preferably 9.5 to 19 wt% so as to further improve the water resistance and prevent the crystal precipitation in the glass in the heat treatment of the glass composition after the manufacture thereof.

Table 4 shows the glass compositions manufactured to make a study of the content of Bi_2O_3 . The glass compositions of Nos. 22 to 27 are Examples and those of Nos. 21 and 28 are Comparative Examples.

Table 4

		Comparative Example	Example						Comparative Example
No.		21	22	23	24	25	26	27	28
Composition	SiO ₂	12.0	11.2	7.0	1.7	2.0	2.1	0.6	0.5
(wt%)	B ₂ O ₃	8.6	8.6	7.9	7.1	5.0	3.7	4.1	3.0
	ZnO	22.0	20.3	16.7	12.1	11.5	7.9	4.5	4.0
	Bi ₂ O ₃	52.5	55.0	62.0	71.0	80.0	85.0	90.0	92.0
	Al ₂ O ₃	1.0	1.0	0.9	0.8	0.5	0.8	—	0.1
	Li ₂ O	0.5	0.5	—	—	—	0.4	—	—
	Na ₂ O	1.5	1.5	1.4	0.3	—	—	0.5	0.2
	K ₂ O	—	—	—	—	—	—	—	—
	MgO	—	—	—	—	—	—	—	—
	CaO	1.0	1.0	0.6	—	—	—	—	—
	Sc ₂ O ₃	—	—	—	1.0	—	—	—	—
	Y ₂ O ₃	—	—	—	1.0	—	—	—	—
	La ₂ O ₃	0.9	0.9	—	1.0	0.5	—	—	—
	CeO ₂	—	—	—	1.0	—	—	—	0.2
	Nd ₂ O ₃	—	—	—	1.0	—	—	—	—
	Gd ₂ O ₃	—	—	3.5	1.0	—	—	0.3	—
	Er ₂ O ₃	—	—	—	1.0	—	0.1	—	—
Stability		○	○	○	○	○	○	△	×
Working Temperature (°C)		660	640	585	525	485	465	450	—
Thermal expansion coefficient (×10 ⁻⁷ /°C)		80	85	92	101	117	110	130	—
Strength (MPa)		140	135	129	122	120	115	110	—
Water resistance (mg/cm ²)		0.0	0.2	0.2	0.3	0.3	0.5	0.8	—

As clearly shown in Table 4, when the Bi₂O₃ content was 52.5 wt%, the working temperature was 660°C, which exceeded 650°C. When the Bi₂O₃ content was 92.0 wt%, the glass composition was devitrified in the manufacture thereof and the glass composition was not obtained with stability. Accordingly, the Bi₂O₃ content is preferably 55 to 90 wt%.

If the glass composition is used particularly as the sealing member of the magnetic head, the Bi_2O_3 content is preferably 62 to 80 wt% so as to further lower the working temperature and prevent the crystal precipitation in the glass in the heat treatment of the glass composition after the manufacture thereof.

Table 5 shows the glass compositions manufactured to make a study of the content of Al_2O_3 . The glass compositions of Nos. 29 to 32 are Examples and the glass composition of No. 33 is Comparative Example.

Table 5

		Example				Comparative Example
No.		29	30	31	32	33
Composition (wt%)	SiO_2	2.0	12.0	4.7	9.0	10.0
	B_2O_3	4.5	7.0	5.5	9.6	9.0
	ZnO	11.5	14.5	12.9	18.3	18.1
	Bi_2O_3	80.5	64.2	73.0	58.2	57.7
	Al_2O_3	—	0.1	2.0	4.0	5.0
	Na_2O	1.0	0.7	1.2	0.8	—
	Y_2O_3	—	—	0.7	—	—
	La_2O_3	0.5	0.5	—	—	—
	CeO_2	—	1.0	—	—	—
	Nd_2O_3	—	—	—	0.1	—
	Gd_2O_3	—	—	—	—	0.2
Stability		\triangle	\bigcirc	\bigcirc	\bigcirc	\bigcirc
Working temperature ($^{\circ}\text{C}$)		470	635	545	650	660
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		122	82	91	77	73
Mechanical strength (MPa)		121	138	130	137	139
Water resistance (mg/cm^2)		0.6	0.0	0.2	0.0	0.0

Though Al_2O_3 is not always an essential component,

it functions to accelerate the vitrification of the glass composition and improve the water resistance. As clearly shown in Table 5, when the Al_2O_3 content was 5 wt%, the working temperature was 660°C , which exceeded 650°C . Therefore, the Al_2O_3 content is preferably 4 wt% or less.

If the glass composition does not contain Al_2O_3 , the crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity. Therefore, if the glass composition is used particularly as the sealing member for the magnetic head, the Al_2O_3 content is preferably 0.1 to 4 wt%. Moreover, in order to further lower the working temperature, it is more preferable to set the Al_2O_3 content to 0.1 to 2 wt%.

Tables 6 and 7 show the glass compositions manufactured to make a study of the contents of Li_2O , Na_2O and K_2O . The glass compositions of Nos. 34 to 44 are Examples and those of Nos. 45 to 48 are Comparative Examples.

Table 6

		Example							
No.		34	35	36	37	38	39	40	41
Composition	SiO ₂	1.2	1.2	1.2	4.0	4.0	4.0	1.2	1.2
(wt%)	B ₂ O ₃	6.7	6.7	6.7	9.0	9.0	9.0	6.7	6.7
	ZnO	7.8	7.8	7.8	13.6	13.6	13.6	7.8	7.8
	Bi ₂ O ₃	80.3	80.3	79.3	67.4	67.4	67.4	80.3	79.3
	Al ₂ O ₃	0.5	0.5	0.5	1.0	1.0	1.0	0.5	0.5
	Li ₂ O	1.0	1.0	1.0	2.0	—	—	3.0	—
	Na ₂ O	1.0	2.0	1.0	1.0	3.0	—	—	4.0
	K ₂ O	1.0	—	2.0	1.0	1.0	4.0	—	—
	La ₂ O ₃	0.5	0.5	0.5	—	—	—	0.5	0.5
	CeO ₂	—	—	—	1.0	1.0	1.0	—	—
Stability		○	○	○	○	○	○	△	△
Working temperature (°C)		475	475	470	585	585	580	470	465
Thermal expansion coefficient (×10 ⁻⁷ /°C)		125	124	127	109	108	112	125	127
Mechanical strength (MPa)		121	120	122	128	127	126	118	119
Water resistance (mg/cm ²)		0.6	0.5	0.7	0.3	0.2	0.4	0.5	0.6

Table 7

		Example			Comparative Example			
No.		42	43	44	45	46	47	48
Composition	SiO ₂	1.2	3.0	3.0	1.0	1.0	1.0	1.0
(wt%)	B ₂ O ₃	6.7	9.0	9.0	6.5	6.5	6.5	6.5
	ZnO	7.8	13.6	13.6	7.5	7.5	7.5	7.5
	Bi ₂ O ₃	78.3	67.4	67.4	77.3	77.3	77.3	77.3
	Al ₂ O ₃	0.5	1.0	1.0	0.2	0.2	0.2	0.2
	Li ₂ O	—	2.0	—	2.0	7.0	—	—
	Na ₂ O	—	3.0	1.0	2.0	—	7.0	—
	K ₂ O	5.0	—	4.0	3.0	—	—	7.0
	La ₂ O ₃	0.5	—	—	0.5	0.5	0.5	0.5
	CeO ₂	—	1.0	1.0	—	—	—	—
Stability		△	△	△	△	△	△	△
Working temperature (°C)		455	570	565	450	450	450	440
Thermal expansion coefficient (×10 ⁻⁷ /°C)		130	111	114	134	135	137	140
Mechanical strength (MPa)		119	126	124	113	114	111	114
Water resistance (mg/cm ²)		0.7	0.4	0.6	1.2	1.5	1.8	1.9

Though Li₂O, Na₂O and K₂O are not always essential components, the working temperature of the glass composition is lowered if at least one or more of them is contained.

As clearly shown in Tables 6 and 7, if the total content of Li₂O, Na₂O and K₂O was 7 wt%, the thermal expansion coefficient exceeded 130×10⁻⁷/°C and a decrease in weight per unit area, which represents the water resistance, also exceeded 1.0 mg/cm². Accordingly, the total content of Li₂O, Na₂O and K₂O is preferably 5 wt% or less, more preferably 4 wt% or less.

When the Li₂O content was 3 wt%, the Na₂O content was 4 wt% and the K₂O content was 5 wt%, the crystal precipitation

was apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof. Accordingly, it is preferable to set the Li_2O content to 2 wt% or less, the Na_2O content to 3 wt% or less and the K_2O content to 4 wt% or less.

Tables 8 and 9 show the glass compositions manufactured to make a study of the contents of MgO , CaO , SrO and BaO . The glass compositions of Nos. 49 to 62 are Examples and those of Nos. 63 to 67 are Comparative Examples.

Table 8

		Example									
No.		49	50	51	52	53	54	55	56	57	58
Composition	SiO_2	2.9	2.6	2.6	2.6	2.6	4.0	2.6	2.6	2.6	2.6
(wt%)	B_2O_3	3.7	6.3	6.3	6.3	6.3	5.0	7.3	7.3	6.3	6.3
	ZnO	5.7	8.0	8.0	8.0	8.0	8.5	9.0	9.0	8.0	8.0
	Bi_2O_3	80.5	70.9	70.9	70.9	70.9	71.5	72.9	71.9	71.9	68.9
	Al_2O_3	1.1	1.0	1.0	1.0	1.0	0.5	1.0	1.0	1.0	1.0
	Na_2O	1.1	0.2	0.2	0.2	0.2	—	0.2	0.2	0.2	0.2
	MgO	1.0	5.0	2.0	—	—	1.0	6.0	—	—	—
	CaO	1.0	1.0	6.0	—	—	2.0	—	7.0	—	—
	SrO	1.0	2.0	2.0	8.0	—	3.0	—	—	9.0	—
	BaO	1.0	2.0	—	2.0	10.0	4.0	—	—	—	12.0
	La_2O_3	0.5	1.0	1.0	1.0	1.0	—	1.0	1.0	1.0	1.0
	CeO_2	0.5	—	—	—	—	0.5	—	—	—	—
Stability		○	○	○	○	○	○	△	△	△	△
Working temperature (°C)		485	530	530	525	525	520	520	525	520	535
Thermal expansion coefficient ($\times 10^{-7}/^\circ\text{C}$)		124	103	103	105	106	102	105	103	103	102
Mechanical strength (MPa)		120	122	122	120	119	128	124	123	122	125
Water resistance (mg/cm^2)		0.4	0.3	0.3	0.4	0.3	0.3	0.2	0.3	0.3	0.3

Table 9

		Example				Comparative Example				
No.		59	60	61	62	63	64	65	66	67
Composition	SiO ₂	4.0	4.0	4.0	4.0	3.0	3.0	3.0	3.0	3.0
(wt%)	B ₂ O ₃	5.0	5.0	5.0	5.0	4.0	4.0	4.0	4.0	4.0
	ZnO	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5
	Bi ₂ O ₃	69.5	69.5	69.5	69.5	69.5	69.5	69.5	69.5	69.5
	Al ₂ O ₃	0.5	0.5	0.5	0.5	0.5	0.5	0.5	0.5	0.5
	MgO	5.0	1.0	—	—	3.0	14.0	—	—	—
	CaO	3.0	6.0	1.0	1.0	3.0	—	14.0	—	—
	SrO	—	2.0	8.0	1.0	4.0	—	—	14.0	—
	BaO	4.0	3.0	3.0	10.0	4.0	—	—	—	14.0
	La ₂ O ₃	—	—	—	—	—	—	—	—	—
	CeO ₂	0.5	0.5	0.5	0.5	0.5	0.5	0.5	0.5	0.5
Stability		△	△	△	△	×	×	×	×	×
Working temperature (°C)		525	520	520	520	—	—	—	—	—
Thermal expansion coefficient (×10 ⁻⁷ /°C)		100	102	102	104	—	—	—	—	—
Mechanical strength (MPa)		122	122	121	123	—	—	—	—	—
Water resistance (mg/cm ²)		0.4	0.3	0.3	0.3	—	—	—	—	—

Though MgO, CaO, SrO and BaO are not always essential components, a glass composition having stability is obtained if at least one or more of them is contained.

As clearly shown in Tables 8 and 9, when the total content of MgO, CaO, SrO and BaO was 14 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass was not obtained with stability. Accordingly, the total content of MgO, CaO, SrO and BaO is preferably 12 wt% or less.

Further, when the MgO content was 6.0 wt%, the CaO content was 7.0 wt%, the SrO content was 9.0 wt% and the BaO

content was 12.0 wt%, the crystal precipitation was apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof. In order to suppress this, it is preferable to set the MgO content to 5 wt% or less, the CaO content to 6 wt% or less, the SrO content to 8 wt% or less and the BaO content to 10 wt% or less, and the total content of them to 10 wt% or less.

Tables 10 to 15 show the glass compositions manufactured to make a study of the contents of Sc_2O_3 , Y_2O_3 , La_2O_3 , CeO_2 , Pr_2O_3 , Nd_2O_3 , Sm_2O_3 , Eu_2O_3 , Gd_2O_3 , Tb_2O_3 , Dy_2O_3 , Ho_2O_3 , Er_2O_3 , Tm_2O_3 , Yb_2O_3 and Lu_2O_3 . The glass compositions of Nos. 69 to 100 are Examples and those of Nos. 68 and 101 to 116 are Comparative Examples.

Table 10

		Comparative Example	Example						
No.		68	69	70	71	72	73	74	75
Composition	SiO ₂	4.8	4.8	4.8	4.8	4.8	4.8	4.8	4.8
(wt%)	B ₂ O ₃	7.6	7.6	7.6	7.6	7.6	7.6	7.6	7.6
	ZnO	12.9	12.9	12.9	12.9	12.9	12.9	12.9	12.9
	Bi ₂ O ₃	73.7	73.6	73.6	73.6	73.6	73.6	73.6	73.6
	Al ₂ O ₃	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0
	Sc ₂ O ₃	—	0.1	—	—	—	—	—	—
	Y ₂ O ₃	—	—	0.1	—	—	—	—	—
	La ₂ O ₃	—	—	—	0.1	—	—	—	—
	CeO ₂	—	—	—	—	0.1	—	—	—
	Pr ₂ O ₃	—	—	—	—	—	0.1	—	—
	Nd ₂ O ₃	—	—	—	—	—	—	0.1	—
	Sm ₂ O ₃	—	—	—	—	—	—	—	0.1
Stability		○	○	○	○	○	○	○	○
Working temperature (°C)		570	570	570	570	570	570	570	570
Thermal expansion coefficient (×10 ⁻⁷ /°C)		88	88	88	87	88	88	88	88
Mechanical strength (MPa)		95	111	112	113	111	110	112	111
Water resistance (mg/cm ²)		0.2	0.2	0.1	0.2	0.2	0.3	0.2	0.2

Table 11

		Example							
No.		76	77	78	79	80	81	82	83
Composition	SiO ₂	4.8	4.8	4.8	4.8	4.8	4.8	4.8	4.8
(wt%)	B ₂ O ₃	7.6	7.6	7.6	7.6	7.6	7.6	7.6	7.6
	ZnO	12.9	12.9	12.9	12.9	12.9	12.9	12.9	12.9
	Bi ₂ O ₃	73.6	73.6	73.6	73.6	73.6	73.6	73.6	73.6
	Al ₂ O ₃	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0
	Eu ₂ O ₃	0.1	—	—	—	—	—	—	—
	Gd ₂ O ₃	—	0.1	—	—	—	—	—	—
	Tb ₂ O ₃	—	—	0.1	—	—	—	—	—
	Dy ₂ O ₃	—	—	—	0.1	—	—	—	—
	Ho ₂ O ₃	—	—	—	—	0.1	—	—	—
	Er ₂ O ₃	—	—	—	—	—	0.1	—	—
	Tm ₂ O ₃	—	—	—	—	—	—	0.1	—
	Yb ₂ O ₃	—	—	—	—	—	—	—	0.1
Stability		○	○	○	○	○	○	○	○
Working temperature (°C)		570	570	570	570	570	570	570	570
Thermal expansion coefficient (×10 ⁻⁷ /°C)		88	88	88	88	88	87	88	88
Mechanical strength (MPa)		111	113	112	110	111	110	112	110
Water resistance (mg/cm ²)		0.2	0.1	0.2	0.2	0.1	0.2	0.2	0.2

Table 12

		Example							
No.		84	85	86	87	88	89	90	91
Composition	SiO ₂	4.8	4.3	4.3	4.3	4.3	4.3	4.3	4.3
(wt%)	B ₂ O ₃	7.6	6.8	6.8	6.8	6.8	6.8	6.8	6.8
	ZnO	12.9	11.6	11.6	11.6	11.6	11.6	11.6	11.6
	Bi ₂ O ₃	73.6	66.3	66.3	66.3	66.3	66.3	66.3	66.3
	Al ₂ O ₃	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0
	Sc ₂ O ₃	—	10.0	—	—	—	—	—	—
	Y ₂ O ₃	—	—	10.0	—	—	—	—	—
	La ₂ O ₃	—	—	—	10.0	—	—	—	—
	CeO ₂	—	—	—	—	10.0	—	—	—
	Pr ₂ O ₃	—	—	—	—	—	10.0	—	—
	Nd ₂ O ₃	—	—	—	—	—	—	10.0	—
	Sm ₂ O ₃	—	—	—	—	—	—	—	10.0
	Lu ₂ O ₃	0.1	—	—	—	—	—	—	—
Stability		○	○	○	○	○	○	○	○
Working temperature (°C)		570	575	575	575	575	575	575	575
Thermal expansion coefficient (×10 ⁻⁷ /°C)		88	89	89	89	89	88	88	88
Mechanical strength (MPa)		111	128	129	128	129	130	128	128
Water resistance (mg/cm ²)		0.2	0.2	0.1	0.2	0.2	0.1	0.2	0.2

Table 13

		Example								
No.		92	93	94	95	96	97	98	99	100
Composition	SiO ₂	4.3	4.3	4.3	4.3	4.3	4.3	4.3	4.3	4.3
(wt%)	B ₂ O ₃	6.8	6.8	6.8	6.8	6.8	6.8	6.8	6.8	6.8
	ZnO	11.6	11.6	11.6	11.6	11.6	11.6	11.6	11.6	11.6
	Bi ₂ O ₃	66.3	66.3	66.3	66.3	66.3	66.3	66.3	66.3	66.3
	Al ₂ O ₃	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0
	Eu ₂ O ₃	10.0	—	—	—	—	—	—	—	—
	Gd ₂ O ₃	—	10.0	—	—	—	—	—	—	—
	Tb ₂ O ₃	—	—	10.0	—	—	—	—	—	—
	Dy ₂ O ₃	—	—	—	10.0	—	—	—	—	—
	Ho ₂ O ₃	—	—	—	—	10.0	—	—	—	—
	Er ₂ O ₃	—	—	—	—	—	10.0	—	—	—
	Tm ₂ O ₃	—	—	—	—	—	—	10.0	—	—
	Yb ₂ O ₃	—	—	—	—	—	—	—	10.0	—
	Lu ₂ O ₃	—	—	—	—	—	—	—	—	10.0
Stability		○	○	○	○	○	○	○	○	○
Working temperature (°C)		575	580	575	580	575	575	580	580	580
Thermal expansion coefficient (×10 ⁻⁷ /°C)		88	87	88	88	88	87	88	88	87
Mechanical strength (MPa)		129	130	128	129	127	128	128	129	127
Water resistance (mg/cm ²)		0.2	0.2	0.2	0.3	0.2	0.2	0.2	0.2	0.2

Table 14

		Comparative Example							
No.		101	102	103	104	105	106	107	108
Composition	SiO ₂	4.3	4.3	4.3	4.3	4.3	4.3	4.3	4.3
(wt%)	B ₂ O ₃	6.8	6.8	6.8	6.8	6.8	6.8	6.8	6.8
	ZnO	11.6	11.6	11.6	11.6	11.6	11.6	11.6	11.6
	Bi ₂ O ₃	64.3	64.3	64.3	64.3	64.3	64.3	64.3	64.3
	Al ₂ O ₃	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0
	Sc ₂ O ₃	12.0	—	—	—	—	—	—	—
	Y ₂ O ₃	—	12.0	—	—	—	—	—	—
	La ₂ O ₃	—	—	12.0	—	—	—	—	—
	CeO ₂	—	—	—	12.0	—	—	—	—
	Pr ₂ O ₃	—	—	—	—	12.0	—	—	—
	Nd ₂ O ₃	—	—	—	—	—	12.0	—	—
	Sm ₂ O ₃	—	—	—	—	—	—	12.0	—
	Eu ₂ O ₃	—	—	—	—	—	—	—	12.0
Stability		×	×	×	×	×	×	×	×
Working temperature (°C)		—	—	—	—	—	—	—	—
Thermal expansion coefficient (×10 ⁻⁷ /°C)		—	—	—	—	—	—	—	—
Mechanical strength (MPa)		—	—	—	—	—	—	—	—
Water resistance (mg/cm ²)		—	—	—	—	—	—	—	—

Table 15

		Comparative Example							
No.		109	110	111	112	113	114	115	116
Composition	SiO ₂	4.3	4.3	4.3	4.3	4.3	4.3	4.3	4.3
(wt%)	B ₂ O ₃	6.8	6.8	6.8	6.8	6.8	6.8	6.8	6.8
	ZnO	11.6	11.6	11.6	11.6	11.6	11.6	11.6	11.6
	Bi ₂ O ₃	64.3	64.3	64.3	64.3	64.3	64.3	64.3	64.3
	Al ₂ O ₃	1.0	1.0	1.0	1.0	1.0	1.0	1.0	1.0
	Gd ₂ O ₃	12.0	—	—	—	—	—	—	—
	Tb ₂ O ₃	—	12.0	—	—	—	—	—	—
	Dy ₂ O ₃	—	—	12.0	—	—	—	—	—
	Ho ₂ O ₃	—	—	—	12.0	—	—	—	—
	Er ₂ O ₃	—	—	—	—	12.0	—	—	—
	Tm ₂ O ₃	—	—	—	—	—	12.0	—	—
	Yb ₂ O ₃	—	—	—	—	—	—	12.0	—
	Lu ₂ O ₃	—	—	—	—	—	—	—	12.0
Stability		×	×	×	×	×	×	×	×
Working temperature (°C)		—	—	—	—	—	—	—	—
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		—	—	—	—	—	—	—	—
Mechanical strength (MPa)		—	—	—	—	—	—	—	—
Water resistance (mg/cm ²)		—	—	—	—	—	—	—	—

Sc₂O₃, Y₂O₃, La₂O₃, CeO₂, Pr₂O₃, Nd₂O₃, Sm₂O₃, Eu₂O₃, Gd₂O₃, Tb₂O₃, Dy₂O₃, Ho₂O₃, Er₂O₃, Tm₂O₃, Yb₂O₃ and Lu₂O₃ function to improve the mechanical strength of the glass composition.

As clearly shown in Tables 10 to 15, if higher mechanical strength is required for the sealing member, it is necessary to use a glass composition containing 0.1 wt% or more of at least one of them. However, if the total content of them was 12 wt% or more, the glass composition was devitrified in the manufacture thereof and the glass

composition was not obtained with stability. Therefore, the total content of them is preferably 0.1 to 10 wt%.

Table 16 shows the glass compositions manufactured to make a study of the content of Gd_2O_3 among the above-listed rare-earth compounds. The glass compositions of Nos. 119 to 122 are Examples and those of Nos. 117 and 118 are Comparative Examples.

Table 16

		Comparative Example		Example			
No.		117	118	119	120	121	122
Composition	SiO_2	2.10	2.10	2.10	2.08	2.06	2.03
(wt%)	B_2O_3	7.30	7.29	7.29	7.24	7.15	7.04
	ZnO	13.27	13.26	13.26	13.17	13.01	12.81
	Bi_2O_3	75.97	75.93	75.90	75.37	74.46	73.31
	Al_2O_3	0.40	0.40	0.40	0.39	0.39	0.38
	Na_2O	0.96	0.96	0.96	0.95	0.94	0.93
	Gd_2O_3	—	0.06	0.10	0.80	2.00	3.50
Stability		○	○	○	○	○	○
Working temperature (°C)		520	520	525	525	530	534
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}C$)		99	99	98	99	98	99
Mechanical strength (MPa)		98	101	108	115	117	120
Water resistance (mg/cm ²)		0.3	0.2	0.3	0.2	0.2	0.2

As clearly shown in Table 16, when the Gd_2O_3 content was 0.1 to 3.5 wt%, high mechanical strength was obtained, thereby the thermal expansion coefficient was reduced.

Table 17 shows the glass compositions manufactured to make a study of the weight ratio of ZnO to B_2O_3 (ZnO/B_2O_3). The glass compositions of Nos. 124 to 126 are Examples and

those of Nos. 123 and 127 are Comparative Examples.

Table 17

		Comparative Example	Example			Comparative Example
No.		123	124	125	126	127
Composition	SiO ₂	4.4	3.1	7.6	3.5	3.5
(wt%)	B ₂ O ₃	6.0	5.5	7.0	6.0	5.0
	ZnO	4.2	4.4	13.3	16.8	17.0
	Bi ₂ O ₃	80.9	81.5	68.7	72.3	73.1
	Al ₂ O ₃	2.0	0.8	1.1	1.0	1.0
	Na ₂ O	2.0	1.2	—	—	—
	MgO	—	3.0	—	—	—
	SrO	—	—	0.9	—	—
	La ₂ O ₃	0.5	0.5	1.4	—	—
	CeO ₂	—	—	—	0.4	0.4
ZnO/B ₂ O ₃		0.7	0.8	1.9	2.8	3.4
Stability		△	○	○	○	△
Working temperature (°C)		520	510	605	545	540
Thermal expansion coefficient (×10 ⁻⁷ /°C)		113	118	86	85	87
Mechanical strength (MPa)		119	114	125	120	116
Water resistance (mg/cm ²)		0.4	0.4	1.0	0.2	0.0

As clearly shown in Table 17, when the weight ratio of ZnO to B₂O₃ is 0.7 or 3.4, the crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity. Accordingly, to obtain a stable glass composition that hardly causes the crystal precipitation, the weight ratio of ZnO to B₂O₃ is preferably controlled to 0.8 to 2.8.

Table 18 shows the glass compositions manufactured to make a study of the weight ratio of Al₂O₃ to SiO₂

($\text{Al}_2\text{O}_3/\text{SiO}_2$). The glass compositions of Nos. 128 to 130 are Examples and those of Nos. 131 and 132 are Comparative Examples.

Table 18

		Example			Comparative Example	
No.		128	129	130	131	132
Composition	SiO_2	3.4	4.0	5.0	4.0	2.0
(wt%)	B_2O_3	6.5	5.5	5.9	5.5	3.4
	ZnO	10.1	8.5	4.7	12.9	9.0
	Bi_2O_3	78.3	72.9	79.0	72.7	82.6
	Al_2O_3	—	1.2	2.5	2.4	1.8
	Li_2O	—	—	—	—	1.1
	Na_2O	1.2	0.4	2.0	1.2	—
	SrO	—	5.5	—	—	—
	La_2O_3	0.5	2.0	—	—	0.1
	CeO_2	—	—	—	0.3	—
	Gd_2O_3	—	—	—	1.0	—
	Er_2O_3	—	—	0.9	—	—
$\text{Al}_2\text{O}_3/\text{SiO}_2$		0.0	0.3	0.5	0.6	0.9
Stability		○	○	○	△	△
Working temperature (°C)		515	520	530	550	475
Thermal expansion coefficient ($\times 10^{-7}/^\circ\text{C}$)		106	104	108	92	106
Mechanical strength (MPa)		115	118	122	123	110
Water resistance (mg/cm^2)		0.3	0.3	0.4	0.2	0.5

As clearly shown in Table 18, when the weight ratio of Al_2O_3 to SiO_2 is 0.6 or more, the crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity. Accordingly, to obtain a stable glass

composition that hardly causes the crystal precipitation, the weight ratio of Al_2O_3 to SiO_2 is preferably controlled to 0.5 or less.

Table 19 shows lead-containing glass compositions manufactured as Comparative Examples 133 to 139 to make a performance comparison with the bismuth glass composition.

Table 19

		Comparative Example						
No.		133	134	135	136	137	138	139
Composition	SiO_2	0.7	1.5	9.9	3.0	16.5	15.2	23.5
(wt%)	B_2O_3	16.0	7.6	12.3	9.0	3.0	8.5	2.3
	ZnO	8.1	8.9	4.5	10.0	0.5	1.2	2.4
	PbO	73.7	78.6	70.4	74.0	79.5	68.5	65.3
	Al_2O_3	0.5	0.4	1.6	1.0	0.5	4.7	3.2
	Na_2O	—	—	1.3	—	—	1.9	3.3
	K_2O	0.7	1.9	—	—	—	—	—
	CaO	0.3	1.1	—	—	—	—	—
	BaO	—	—	—	0.3	—	—	—
Stability		○	○	○	○	○	○	○
Working temperature ($^{\circ}\text{C}$)		480	480	540	490	530	630	650
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		95	107	91	96	104	88	91
Mechanical strength (MPa)		115	111	127	118	135	130	145
Water resistance (mg/cm^2)		3.3	4.0	1.9	2.8	2.1	1.5	1.3

As clearly shown in Table 19, the lead-containing glasses, in every composition, showed the water resistance higher than $1.0 \text{ mg}/\text{cm}^2$. On the other hand, the bismuth glass composition of the present invention was lead free and showed the working temperature of 450 to 650°C , the thermal expansion coefficient of 70×10^{-7} to $130 \times 10^{-7}/^{\circ}\text{C}$, mechanical strength equal to that of the lead-containing glass composition and

excellent water resistance.

If the glass composition of the present invention is used as the sealing member for the magnetic head, the glass composition is preferably amorphous from the viewpoint of exhibiting flowability. However, depending on the applications of the glass composition, the glass composition may be crystallized by heat treatment.

Table 20 shows more preferable examples of the compositions of the bismuth glass composition of the present invention.

Table 20

		Example				
No.		140	141	142	143	144
Composition (wt%)	SiO ₂	3.4	2.1	8.1	2.1	2.0
	B ₂ O ₃	6.7	7.2	5.7	7.2	7.0
	ZnO	12.5	13.2	13.3	13.2	12.8
	Bi ₂ O ₃	75.7	75.5	68.7	75.3	73.2
	Al ₂ O ₃	1.0	0.4	0.8	0.4	0.4
	Na ₂ O	0.6	1.0	0.3	0.7	0.7
	CaO	—	—	1.2	0.4	0.4
	SrO	—	—	1.2	—	—
	La ₂ O ₃	0.1	0.6	—	—	—
	Nd ₂ O ₃	—	—	0.7	—	—
	Gd ₂ O ₃	—	—	—	0.7	3.5
Stability		○	○	○	○	○
Working temperature (°C)		539	524	605	525	534
Thermal expansion coefficient (×10 ⁻⁷ /°C)		97	98	87	99	99
Mechanical strength (MPa)		126	120	138	123	130
Water resistance (mg/cm ²)		0.3	0.4	0.1	0.3	0.2

The above-described glass compositions may be used

in the form of bulk, powder, fiber or thin film for the manufacture of the magnetic head. However, the form of the glass composition is not limited thereto.

The glass composition or the sealing member for the magnetic head according to the present invention may be used as a material consisted of the glass composition only, or a composite material comprising a combination of the glass composition and other materials. The glass composition or the sealing member for the magnetic head may be crystallized by heat treatment or combined with other materials to be used as a composite material such that it can be used as a material having a thermal expansion coefficient higher than 70×10^{-7} to $130 \times 10^{-7}/^{\circ}\text{C}$.

The temperature for heat-treating the glass composition or the sealing member for the magnetic head according to the present invention is not limited to the working temperature of 450 to 650°C . It is needless to say that the temperature may be selected depending on the applications.

Example 2

As an example of the present invention, a ferrite head shown in FIG. 4 was formed, with which a magnetic recording/reproducing device was manufactured.

Referring to FIG. 4, the glass composition of Example 17 was used as the sealing members 25 and 26, which

were subjected to heat treatment at 610°C to form a ferrite head 20. As ferrite for forming the magnetic core halves 22 and 23, Mn-Zn monocrystalline ferrite was used and quartz glass was used as the magnetic gap member 24.

The thus formed ferrite head was not cracked or broken and the sealing member was not eroded, exhibiting intended magnetic conversion properties. Further, the magnetic recording/reproducing device using the ferrite head was able to perform recording and reproducing with high reliability.

As a sealing member best for the ferrite head, the following glass composition is preferably used.

As the sealing member for the ferrite head, those mentioned as the above Examples may be used. In particular, due to a demand for high wear resistance to a magnetic recording medium, it is more preferable to set the SiO₂ content to 7 wt% or more.

In order to obtain a more stable sealing member for the magnetic head under these conditions, it is necessary to further adjust the content of each component. It is preferred that the composition comprises 7 to 12 wt% of SiO₂, 3 to 9 wt% of B₂O₃, 9.5 to 19 wt% or ZnO, 62 to 80 wt% of Bi₂O₃, 0.1 to 4 wt% of Al₂O₃, 0.1 to 5 wt% of at least one selected from the group consisting of Sc₂O₃, Y₂O₃, La₂O₃, CeO₂, Pr₂O₃, Nd₂O₃, Sm₂O₃, Eu₂O₃, Gd₂O₃, Tb₂O₃, Dy₂O₃, Ho₂O₃, Er₂O₃, Tm₂O₃, Yb₂O₃ and Lu₂O₃.

Further, in order to make an adjustment to obtain

desired working temperature and thermal expansion coefficient, it is preferable to set the contents of Li_2O to 2 wt% or less, Na_2O to 3 wt% or less, K_2O to 4 wt% or less, and the total content of them to 4 wt% or less. In addition, it is preferable to set the contents of MgO to 5 wt% or less, CaO to 6 wt% or less, SrO to 8 wt% or less, BaO to 10 wt% or less, and the total content of them to 10 wt% or less.

Moreover, in the above-described glass composition, the weight ratio of ZnO to B_2O_3 is preferably 0.8 to 2.8 and the weight ratio of Al_2O_3 to SiO_2 is preferably 0.5 or less.

Example 3

As an example of the present invention, an MIG head shown in FIG. 5 was formed, with which a magnetic recording/reproducing device was manufactured.

Referring to FIG. 5, the glass composition of Example 71 was used as the sealing members 35 and 36, which were subjected to heat treatment at 525°C to form an MIG head 30. As ferrite for forming the magnetic core halves 32 and 33, Mn-Zn monocrystalline ferrite was used. Fe-Ta-N films having saturated magnetic flux density (B_s) of 1.6 T were used as the magnetic metal films 37 and 38. Quartz glass was used as the magnetic gap member 34.

The thus formed MIG head was not cracked or broken and the sealing member was not eroded, exhibiting intended magnetic conversion properties. Further, the magnetic

recording/reproducing device using the MIG head was able to perform recording and reproducing with high reliability.

As a sealing member best for the MIG head, the following glass composition is preferably used.

As the sealing member for the MIG head, those mentioned as the above Examples may be used. In particular, for the purpose of manufacturing the magnetic head in high yield, the sealing member is required to have high wettability to the magnetic core halves. To meet this requirement, it is more preferable to set the content of SiO_2 to 7 wt% or less. Further, in order to keep the wear resistance to the magnetic recording medium high, it is more preferable to set the content of SiO_2 to 1.7 wt% or more.

In order to obtain a more stable sealing member for the magnetic head under these conditions, it is necessary to further adjust the content of each component. It is preferred that the composition comprises 1.7 to 7 wt% of SiO_2 , 3 to 9 wt% of B_2O_3 , 9.5 to 19 wt% or ZnO , 68 to 80 wt% of Bi_2O_3 , 0.1 to 4 wt% of Al_2O_3 , 0.1 to 5 wt% of at least one selected from the group consisting of Sc_2O_3 , Y_2O_3 , La_2O_3 , CeO_2 , Pr_2O_3 , Nd_2O_3 , Sm_2O_3 , Eu_2O_3 , Gd_2O_3 , Tb_2O_3 , Dy_2O_3 , Ho_2O_3 , Er_2O_3 , Tm_2O_3 , Yb_2O_3 and Lu_2O_3 .

Further, in order to make an adjustment to obtain desired working temperature and thermal expansion coefficient, it is preferable to set the contents of Li_2O to 2 wt% or less, Na_2O to 3 wt% or less, K_2O to 4 wt% or less, and the total

content of them to 4 wt% or less. In addition, it is preferable to set the contents of MgO to 5 wt% or less, CaO to 6 wt% or less, SrO to 8 wt% or less, BaO to 10 wt% or less, and the total content of them to 10 wt% or less.

Moreover, in the above-described glass composition, the weight ratio of ZnO to B₂O₃ is preferably 0.8 to 2.8 and the weight ratio of Al₂O₃ to SiO₂ is preferably 0.5 or less.

Example 4

As an example of the present invention, a stacked head shown in FIG. 6 was formed, with which a magnetic recording/reproducing device was manufactured.

Referring to FIG. 6, the glass composition of Example 141 was used as the sealing members 45, 46 and 47, which were subjected to heat treatment at 490°C to form a stacked head 40. A Co-Ta-Zr-Nb alloy having saturated magnetic flux density (B_s) of 0.8 T was used as the magnetic metal film 53. Quartz glass was used as the insulating film 54, a ceramic substrate comprising MgO-NiO-TiO₂ was used as the nonmagnetic substrates 51 and 52, and quartz glass was used as the magnetic gap member 44.

The thus formed stacked head was not cracked or broken and the sealing member was not eroded, exhibiting intended magnetic conversion properties. Further, the magnetic recording/reproducing device using the stacked head was able to perform recording and reproducing with high

reliability.

As a sealing member best for the stacked head, the following glass composition is preferably used.

As the sealing member for the stacked head, those mentioned as the above Examples may be used. In particular, in order to prevent deterioration of other members in the heat treatment, it is required to form the magnetic head by heat treatment at a low temperature not higher than 500°C. To meet the requirement, it is preferable to set the content of SiO₂ to 5 wt% or less. Further, to maintain the wear resistance to the magnetic recording medium, it is more preferable to set the content of SiO₂ to 1.7 wt% or more.

In order to obtain a more stable sealing member for the magnetic head under these conditions, it is necessary to further adjust the content of each component. It is preferred that the composition comprises 1.7 to 5 wt% of SiO₂, 3 to 9 wt% of B₂O₃, 9.5 to 19 wt% or ZnO, 68 to 80 wt% of Bi₂O₃, 0.1 to 4 wt% of Al₂O₃, 0.1 to 5 wt% of at least one selected from the group consisting of Sc₂O₃, Y₂O₃, La₂O₃, CeO₂, Pr₂O₃, Nd₂O₃, Sm₂O₃, Eu₂O₃, Gd₂O₃, Tb₂O₃, Dy₂O₃, Ho₂O₃, Er₂O₃, Tm₂O₃, Yb₂O₃ and Lu₂O₃.

Further, in order to make an adjustment to obtain desired working temperature and thermal expansion coefficient, it is preferable to set the contents of Li₂O to 2 wt% or less, Na₂O to 3 wt% or less, K₂O to 4 wt% or less, and the total content of them to 4 wt% or less. Further, it is preferable

to set the contents of MgO to 5 wt% or less, CaO to 6 wt% or less, SrO to 8 wt% or less, BaO to 10 wt% or less, and the total content of them to 10 wt% or less.

Moreover, in the above-described glass composition, the weight ratio of ZnO to B₂O₃ is preferably 0.8 to 2.8 and the weight ratio of Al₂O₃ to SiO₂ is preferably 0.5 or less.

Example 5

Tables 21 to 39 show Examples of the bismuth glass composition according to the present invention used as the sealing member for a PDP, as well as Comparative Examples.

After the predetermined materials were mixed, the obtained mixture was placed in a platinum crucible and melted in an electric furnace at 900 to 1100°C for 1 hour. Then, the obtained molten glass was quenched with a roller to form a glass composition. Tables 21 to 39 show the composition, glass stability, softening point, thermal expansion coefficient, water resistance and strength of the obtained glass compositions.

The glass stability, softening point, thermal expansion coefficient, water resistance and strength were evaluated in the following manner.

[Stability]

After the manufacture of the glass composition, the glass stability was judged from whether or not precipitation of crystals was caused from the glassy state in heat treatment

performed at a temperature higher than each softening point by 50°C for 30 minutes. The stability was indicated by ○ when the glass composition was not devitrified in the manufacture thereof and did not cause the crystal precipitation in the heat treatment at the above-described test temperature, △ when the glass composition was not devitrified in the manufacture thereof but caused the crystal precipitation in the heat treatment at the above-described test temperature, or × when the glass composition was devitrified in the manufacture thereof and the glass composition was not obtained. The glass composition for the sealing member for a PDP may be crystallized after the heat treatment. However, if it is necessary to carry out the heat treatment several times depending on the process, it is preferable that the glass composition does not change in physical property due to the crystallization. It is more preferable that a stable amorphous phase is maintained.

[Softening point]

The softening point was determined in conformity with a test method of JIS R3103-1. More specifically, the glass composition was re-fused to be a melt having a suitable viscosity, from which a glass fiber of 0.55 to 0.75 mm in diameter and 234 to 236 mm in length was drawn. In an oven in which the glass fiber was suspended, temperature was raised to determine the temperature at which the glass fiber was stretched by its own weight at a rate of 1 mm per minute.

In a step of sealing a PDP, if the heat treatment is carried out at around a temperature higher than the softening point of the glass composition by 50°C, the glass composition in the sealing member is enough liquidized and the front plate and the rear plate are well bonded. Accordingly, to perform the sealing at a treatment temperature not higher than 500°C, a glass composition having a softening point of 450°C or less is preferably selected.

[Thermal expansion coefficient]

The glass composition was re-fused, poured into a mold and cooled to form a glass rod of 4 mm in diameter and 20 mm in length. Under heating at 10°C/min, linear expansivity of the glass rod was measured. Then, an average thermal expansion coefficient at 30 to 300°C was calculated.

[Water resistance]

A glass composition having a particle diameter of 250 to 425 μm , equivalent to apparent volume of 1cm^3 , was immersed into boiling ion-exchanged water for an hour. The water resistance was then evaluated from a rate of decrease in weight of the glass composition (%). The smaller the rate of decrease in weight is, the more excellent the water resistance is. Regarding a PDP, 1.5 wt% or less is preferable.

[Mechanical strength]

Mechanical strength was measured by a three-point bending method. The glass composition was re-fused to be a melt having a suitable viscosity, from which a glass rod of 1

mm in diameter and 30 mm in length was drawn. The glass rod was held horizontally by two points at an interval of 20 mm at the center portion thereof, and a load was applied to the center between the two points from above by a load cell at a rate of 1 mm/min. Based on the load that broke the glass rod, the mechanical strength was calculated.

Table 21 shows the glass compositions manufactured to make a study of the content of SiO_2 . The glass compositions of Nos. 203 to 206 are Examples and those of Nos. 201, 202 and 207 are Comparative Examples.

Table 21

		Comparative Example		Example				Comparative Example
No.		201	202	203	204	204	206	207
Composition (wt%)	SiO_2	0.3	0.8	1.1	3.4	4.0	4.5	4.9
	B_2O_3	5.7	8.0	6.4	6.5	4.7	5.9	9.8
	ZnO	10.3	9.5	12.8	10.6	11.6	11.9	9.9
	Bi_2O_3	80.1	79.2	77.6	78.3	79.0	76.2	75.4
	Al_2O_3	0.1	—	0.1	—	0.1	0.5	—
	Na_2O	3.5	0.5	2.0	1.2	0.6	1.0	—
	MgO	—	2.0	—	—	—	—	—
Stability		×	×	○	○	○	△	△
Softening point (°C)		—	—	410	435	444	450	486
Thermal expansion coefficient ($\times 10^{-7}/^\circ\text{C}$)		—	—	108	106	100	100	94
Water resistance (%)		—	—	0.4	0.2	0.2	0.1	0.1
Mechanical strength (MPa)		—	—	72	78	81	80	81

As clearly shown in Table 21, when the SiO_2 content was 0.3 wt% or 0.8 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was

not obtained with stability. When the SiO_2 content was 4.9 wt%, the softening point was 486°C , which exceeded 450°C . From the above results, the SiO_2 content is preferably 1.1 to 4.5 wt%.

Further, when the SiO_2 content is 4.5 wt%, crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may cause a problem in maintaining hermeticity. Therefore, in order to reduce the crystal precipitation in the glass, the SiO_2 content is preferably 1.1 to 4 wt%.

Table 22 shows the glass compositions manufactured to make a study of the content of B_2O_3 . The glass compositions of Nos. 209 to 211 are Examples and those of Nos. 208 and 212 are Comparative Examples.

Table 22

		Comparative Example	Example			Comparative Example
No.		208	209	210	211	212
Composition (wt%)	SiO ₂	4.0	3.4	1.2	2.7	4.4
	B ₂ O ₃	2.3	4.0	7.4	9.0	10.6
	ZnO	16.2	10.0	16.1	9.9	10.6
	Bi ₂ O ₃	77.5	76.2	73.7	77.1	72.4
	Al ₂ O ₃	—	1.0	0.2	0.2	—
	Li ₂ O	—	1.5	—	—	—
	Na ₂ O	—	—	1.2	1.1	—
	K ₂ O	—	—	—	—	2.0
	SrO	—	3.9	—	—	—
	La ₂ O ₃	—	—	0.2	—	—
Stability		×	○	○	○	×
Softening point (°C)		—	407	433	449	—
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		—	110	103	104	—
Water resistance (%)		—	0.3	0.2	0.3	—
Mechanical strength (MPa)		—	80	79	81	—

As clearly shown in Table 22, when the B₂O₃ content was 2.3 wt% or 10.6 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was not obtained with stability. Therefore, the B₂O₃ content is preferably 4 to 9 wt%.

Table 23 shows the glass compositions manufactured to make a study of the content of ZnO. The glass compositions of Nos. 214 to 216 are Examples and those of Nos. 213 and 217 are Comparative Examples.

Table 23

		Comparative Example	Example				Comparative Example
No.		213	214	215	216	217	
Composition (wt%)	SiO ₂	2.1	3.3	1.1	1.2	3.1	
	B ₂ O ₃	6.2	5.8	7.1	7.3	4.5	
	ZnO	7.6	9.5	13.5	18.0	19.2	
	Bi ₂ O ₃	81.7	80.0	76.1	72.8	72.0	
	Al ₂ O ₃	0.9	0.9	0.2	0.1	0.5	
	Li ₂ O	—	—	0.9	—	—	
	Na ₂ O	—	0.5	—	0.6	0.7	
	K ₂ O	0.5	—	—	—	—	
	MgO	1.0	—	—	—	—	
	SrO	—	—	1.0	—	—	
	La ₂ O ₃	—	—	0.1	—	—	
Stability		○	○	○	○	×	
Softening point (°C)		440	448	413	437	—	
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		103	101	105	97	—	
Water resistance (%)		1.6	0.5	0.3	0.2	—	
Mechanical strength (MPa)		81	82	77	79	—	

As clearly shown in Table 23, when the ZnO content was 7.6 wt%, the weight decrease rate representing the water resistance was 1.6 %, which exceeded 1.5 %. When the ZnO content was 19.2 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was not obtained with stability. Therefore, the ZnO content is preferably 9.5 to 18 wt%.

Table 24 shows the glass compositions manufactured to make a study of the content of Bi₂O₃. The glass compositions of Nos. 219 to 222 are Examples and those of Nos. 218 and 223 are Comparative Examples.

Table 24

		Comparative Example	Example					Comparative Example
No.		218	219	220	221	222	223	
Composition (wt%)	SiO ₂	4.5	3.5	1.5	2.0	1.5	3.5	
	B ₂ O ₃	7.6	6.3	6.1	5.5	4.0	2.0	
	ZnO	13.0	10.9	11.7	9.5	9.5	7.1	
	Bi ₂ O ₃	70.4	72.0	79.5	82.0	85.0	87.4	
	Al ₂ O ₃	1.0	1.1	0.1	0.8	—	—	
	Na ₂ O	—	1.0	—	0.2	—	—	
	K ₂ O	0.4	1.2	—	—	—	—	
	MgO	1.6	—	—	—	—	—	
	CaO	—	—	1.0	—	—	—	
	SrO	1.5	—	—	—	—	—	
	BaO	—	4.0	—	—	—	—	
	La ₂ O ₃	—	—	0.1	—	—	—	
Stability		○	○	○	○	△	×	
Softening point (°C)		491	446	441	433	415	—	
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		92	105	101	105	106	—	
Water resistance (%)		0.1	0.2	0.1	0.3	0.4	—	
Mechanical strength (MPa)		85	81	78	77	74	—	

As clearly shown in Table 24, when the Bi₂O₃ content was 70.4 wt%, the softening point was 491°C, which exceeded 450°C. When the Bi₂O₃ content was 87.4 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was not obtained with stability. Therefore, the Bi₂O₃ content is preferably 72 to 85 wt%.

Further, when the Bi₂O₃ content is 85.0 wt%, crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may cause a problem in maintaining hermeticity.

Therefore, in order to reduce the crystal precipitation in the glass, the Bi_2O_3 content is preferably 72 to 82 wt%.

Table 25 shows the glass compositions manufactured to make a study of the content of Al_2O_3 . The glass compositions of Nos. 224 to 227 are Examples and the glass composition of No. 228 is Comparative Example.

Table 25

		Example				Comparative Example
No.		224	225	226	227	228
Composition (wt%)	SiO_2	3.4	1.8	1.1	4.0	4.3
	B_2O_3	6.5	7.0	5.1	5.1	5.8
	ZnO	10.6	13.1	11.6	10.9	11.8
	Bi_2O_3	78.3	76.0	81.0	76.8	72.2
	Al_2O_3	—	0.1	0.2	2.0	2.7
	Li_2O	—	0.1	—	—	—
	Na_2O	1.2	0.7	—	1.0	0.6
	CaO	—	0.4	—	0.2	0.9
	SrO	—	0.8	0.9	—	1.7
	La_2O_3	—	—	0.1	—	—
Stability		○	○	○	○	○
Softening point (°C)		435	436	420	450	471
Thermal expansion coefficient ($\times 10^{-7}/^\circ\text{C}$)		106	104	103	100	97
Water resistance (%)		0.2	0.3	0.2	0.1	0.1
Mechanical strength (MPa)		78	77	75	80	82

Though Al_2O_3 is not always an essential component, it functions to accelerate the vitrification of the glass composition and improve the water resistance. As clearly shown in Table 25, when the Al_2O_3 content was 2.7 wt%, the softening point was 471°C , which exceeded 450°C . Therefore, the Al_2O_3 content is preferably 2 wt% or less.

Further, in order to keep the glassy state stable and to reduce the crystals precipitated in the glass in the heat treatment after the manufacture thereof, the Al_2O_3 content is preferably 0.1 wt% or more.

Tables 26 and 27 show the glass compositions manufactured to make a study of the contents of Li_2O , Na_2O and K_2O . The glass compositions of Nos. 229 to 237 are Examples and those of Nos. 238 to 242 are Comparative Examples.

Table 26

		Example						
No.		229	230	231	232	233	234	235
Composition	SiO_2	1.9	1.9	1.9	1.9	1.9	1.9	1.9
(wt%)	B_2O_3	6.9	6.9	6.9	6.9	6.9	6.9	6.9
	ZnO	12.0	12.0	12.0	12.0	12.0	12.0	12.0
	Bi_2O_3	79.0	77.0	77.0	77.0	75.0	75.0	75.0
	Al_2O_3	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Li_2O	—	2.0	—	—	2.0	1.0	4.0
	Na_2O	—	—	2.0	—	—	2.0	—
	K_2O	—	—	—	2.0	2.0	1.0	—
Stability		○	○	○	○	○	○	△
Softening point (°C)		445	381	419	431	372	387	369
Thermal expansion Coefficient ($\times 10^{-7}/^\circ\text{C}$)		97	112	112	105	119	123	126
Water resistance (%)		0.1	0.3	0.3	0.3	0.3	0.4	0.5
Mechanical strength (MPa)		80	75	77	74	69	70	68

Table 27

		Example		Comparative Example				
No.		236	237	238	239	240	241	242
Composition	SiO ₂	1.9	1.9	1.9	1.9	1.9	1.9	1.9
(wt%)	B ₂ O ₃	6.9	6.9	6.9	6.9	6.9	6.9	6.9
	ZnO	12.0	12.0	12.0	12.0	12.0	12.0	12.0
	Bi ₂ O ₃	75.0	75.0	74.0	74.0	74.0	74.0	74.0
	Al ₂ O ₃	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Li ₂ O	—	—	2.0	1.0	5.0	—	—
	Na ₂ O	4.0	—	2.0	2.0	—	5.0	—
	K ₂ O	—	4.0	1.0	2.0	—	—	5.0
Stability		○	△	△	△	△	△	△
Softening point (°C)		394	438	360	385	360	381	437
Thermal expansion Coefficient (×10 ⁻⁷ /°C)		126	111	130	126	134	133	114
Water resistance (%)		0.4	0.5	1.6	1.6	1.7	1.7	1.8
Mechanical strength (MPa)		71	76	68	70	67	69	78

Though Li₂O, Na₂O and K₂O are not always essential components, the softening point of the glass composition is lowered if at least one or more of them is contained. As clearly shown in Comparative Examples Nos. 238 to 242, if the total content of Li₂O, Na₂O and K₂O was 5.0 wt%, the water resistance was 1.6 to 1.8 %, which exceeded 1.5 %. Therefore, the total content of Li₂O, Na₂O and K₂O is preferably 4 wt% or less.

When the Li₂O content is 4.0 wt% and the K₂O content is 4.0 wt%, the crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity. Therefore, in order to reduce the

crystal precipitation in the glass, it is more preferable that the contents of Li_2O and K_2O are 2 wt% or less, respectively.

In particular, Na_2O is the most effective for maintaining the glass composition stable, though the softening point is low.

However, if the insulating properties are concerned, it is preferable to contain Li_2O , Na_2O and K_2O in an amount as small as possible or not to contain them.

Tables 28 to 30 show the glass compositions manufactured to make a study of the contents of MgO , CaO , SrO and BaO . The glass compositions of Nos. 243 to 261 are Examples and those of Nos. 262 and 266 are Comparative Examples.

Table 28

		Example							
No.		243	244	245	246	247	248	249	250
Composition (wt%)	SiO ₂	1.3	1.3	1.3	1.3	1.3	1.3	1.3	1.3
	B ₂ O ₃	5.3	5.3	5.3	5.3	5.3	5.3	5.3	5.3
	ZnO	10.0	10.0	10.0	10.0	10.0	10.0	10.0	10.0
	Bi ₂ O ₃	83.0	82.9	82.9	81.0	78.5	78.5	79.0	78.5
	Al ₂ O ₃	0.4	0.4	0.4	0.4	0.4	0.4	0.4	0.4
	MgO	—	—	—	2.0	—	—	—	2.0
	CaO	—	0.1	—	—	4.5	—	—	1.0
	SrO	—	—	0.1	—	—	4.5	—	1.5
	BaO	—	—	—	—	—	—	4.0	—
Stability		○	○	○	○	○	○	○	○
Softening point (°C)		430	430	430	444	449	437	431	450
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		102	102	102	102	108	105	104	104
Water resistance (%)		0.1	0.1	0.1	0.2	0.2	0.2	0.3	0.2
Mechanical strength (MPa)		82	87	86	89	90	91	88	90

Table 29

		Example							
No.		251	252	253	254	255	256	257	258
Composition (wt%)	SiO ₂	1.3	1.3	1.3	1.3	1.3	1.3	1.3	1.3
	B ₂ O ₃	5.3	5.3	5.3	5.3	5.3	5.3	5.3	5.3
	ZnO	10.0	10.0	10.0	10.0	10.0	10.0	10.0	10.0
	Bi ₂ O ₃	78.5	78.5	78.5	78.5	77.5	75.0	75.0	75.0
	Al ₂ O ₃	0.4	0.4	0.4	0.4	0.4	0.4	0.4	0.4
	MgO	—	3.0	—	—	1.0	2.0	—	8.0
	CaO	1.5	—	1.5	—	1.0	3.0	—	—
	SrO	1.0	1.5	—	—	1.5	3.0	4.0	—
	BaO	2.0	—	3.0	4.5	2.0	—	4.0	—
Stability		○	△	○	△	△	△	△	△
Softening point (°C)		439	450	437	431	445	449	438	450
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		106	103	106	104	105	108	107	113.0
Water resistance (%)		0.3	0.3	0.5	0.7	1.6	1.6	1.6	1.6
Mechanical strength (MPa)		88	87	88	87	87	91	90	88

Table 30

		Example			Comparative Example				
No.		259	260	261	262	263	264	265	266
Composition (wt%)	SiO ₂	1.3	1.3	1.3	1.3	1.3	1.3	1.3	1.3
	B ₂ O ₃	5.3	5.3	5.3	5.3	5.3	5.3	5.3	5.3
	ZnO	10.0	10.0	10.0	10.0	10.0	10.0	10.0	10.0
	Bi ₂ O ₃	75.0	75.0	75.0	73.0	73.0	73.0	73.0	73.0
	Al ₂ O ₃	0.4	0.4	0.4	0.4	0.4	0.4	0.4	0.4
	MgO	—	—	—	1.0	10.0	—	—	—
	CaO	8.0	—	—	3.0	—	10.0	—	—
	SrO	—	8.0	—	3.0	—	—	10.0	—
	BaO	—	—	8.0	3.0	—	—	—	10.0
Stability		△	△	△	×	×	×	×	×
Softening point (°C)		449	443	433	—	—	—	—	—
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)		113.0	108.0	107.0	—	—	—	—	—
Water resistance (%)		1.6	1.6	1.7	—	—	—	—	—
Mechanical strength (MPa)		90	90	86	—	—	—	—	—

Though MgO, CaO, SrO and BaO are not always essential components, stable glass is formed if one or more of them is contained. As clearly shown in Comparative Examples Nos. 262 to 266, when the total content of MgO, CaO, SrO and BaO was 10 wt%, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was not obtained with stability. Accordingly, the total content of MgO, CaO, SrO and BaO is preferably 8 wt% or less.

Further, as clearly understood from Examples Nos. 255 to 261, when the total content of MgO, CaO, SrO and BaO was 5.5 wt% to 8.0 wt%, the water resistance was 1.6 % or more, which exceeded 1.5 %. For the purpose of improving the water resistance of the glass composition, the total content of MgO,

CaO, SrO and BaO is preferably 4.5 wt% or less.

Further, as clearly shown in Examples Nos. 252 and 254, crystal precipitation is apt to occur in the glass in the heat treatment after the manufacture of the glass composition. In order to suppress this, it is more preferable to set the content of MgO to 2 wt% or less and the content of BaO to 4 wt% or less.

Any of MgO, CaO, SrO and BaO functions to stabilize the glass composition. In particular, it is preferable to contain CaO or SrO, and the effective content thereof is 0.1 to 4.5 wt%, respectively.

Tables 31 to 36 shows the glass compositions manufactured to make a study of the contents of Sc_2O_3 , Y_2O_3 , La_2O_3 , CeO_2 , Pr_2O_3 , Nd_2O_3 , Sm_2O_3 , Eu_2O_3 , Gd_2O_3 , Tb_2O_3 , Dy_2O_3 , Ho_2O_3 , Er_2O_3 , Tm_2O_3 , Yb_2O_3 and Lu_2O_3 . The glass compositions of Nos. 268 to 300 are Examples and those of Nos. 267 and 301 to 317 are Comparative Examples.

Table 31

		Comparative Example	Example							
No.		267	268	269	270	271	272	273	274	275
Composition	SiO ₂	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1
(wt%)	B ₂ O ₃	7.0	7.0	7.0	7.0	7.0	7.0	7.0	7.0	7.0
	ZnO	10.8	10.8	10.8	10.8	10.8	10.8	10.8	10.8	10.8
	Bi ₂ O ₃	79.1	79.0	79.0	79.0	79.0	79.0	79.0	79.0	79.0
	Al ₂ O ₃	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Na ₂ O	1.8	1.8	1.8	1.8	1.8	1.8	1.8	1.8	1.8
	Sc ₂ O ₃	—	0.1	—	—	—	—	—	—	—
	Y ₂ O ₃	—	—	0.1	—	—	—	—	—	—
	La ₂ O ₃	—	—	—	0.1	—	—	—	—	—
	CeO ₂	—	—	—	—	0.1	—	—	—	—
	Pr ₂ O ₃	—	—	—	—	—	0.1	—	—	—
	Nd ₂ O ₃	—	—	—	—	—	—	0.1	—	—
	Sm ₂ O ₃	—	—	—	—	—	—	—	0.1	—
	Eu ₂ O ₃	—	—	—	—	—	—	—	—	0.1
Stability		○	○	○	○	○	○	○	○	○
Softening point (°C)		414	413	414	413	415	413	413	412	415
Thermal expansion coefficient (×10 ⁻⁷ /°C)		113	112	113	113	113	112	113	114	113
Water resistance (%)		0.3	0.2	0.3	0.3	0.3	0.3	0.2	0.2	0.3
Mechanical strength (MPa)		73	82	83	82	85	84	82	81	83

Table 32

		Example							
No.		276	277	278	279	280	281	282	283
Composition	SiO ₂	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1
(wt%)	B ₂ O ₃	7.0	7.0	7.0	7.0	7.0	7.0	7.0	7.0
	ZnO	10.8	10.8	10.8	10.8	10.8	10.8	10.8	10.8
	Bi ₂ O ₃	79.0	79.0	79.0	79.0	79.0	79.0	79.0	79.0
	Al ₂ O ₃	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Na ₂ O	1.8	1.8	1.8	1.8	1.8	1.8	1.8	1.8
	Gd ₂ O ₃	0.1	—	—	—	—	—	—	—
	Tb ₂ O ₃	—	0.1	—	—	—	—	—	—
	Dy ₂ O ₃	—	—	0.1	—	—	—	—	—
	Ho ₂ O ₃	—	—	—	0.1	—	—	—	—
	Er ₂ O ₃	—	—	—	—	0.1	—	—	—
	Tm ₂ O ₃	—	—	—	—	—	0.1	—	—
	Yb ₂ O ₃	—	—	—	—	—	—	0.1	—
	Lu ₂ O ₃	—	—	—	—	—	—	—	0.1
Stability		○	○	○	○	○	○	○	○
Softening point (°C)		415	413	413	412	415	414	413	412
Thermal expansion coefficient (×10 ⁻⁷ /°C)		113	113	113	112	113	112	114	113
Water resistance (%)		0.3	0.2	0.3	0.3	0.3	0.2	0.3	0.3
Mechanical strength (MPa)		84	82	83	81	85	84	83	81

Table 33

		Example								
No.		284	285	286	287	288	289	290	291	292
Composition	SiO ₂	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1
(wt%)	B ₂ O ₃	6.7	6.6	6.6	6.6	6.6	6.6	6.6	6.6	6.6
	ZnO	10.5	10.3	10.3	10.3	10.3	10.3	10.3	10.3	10.3
	Bi ₂ O ₃	76.6	75.1	75.1	75.1	75.1	75.1	75.1	75.1	75.1
	Al ₂ O ₃	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Na ₂ O	1.7	1.7	1.7	1.7	1.7	1.7	1.7	1.7	1.7
	Sc ₂ O ₃	0.2	5.0	—	—	—	—	—	—	—
	Y ₂ O ₃	0.2	—	5.0	—	—	—	—	—	—
	La ₂ O ₃	0.2	—	—	5.0	—	—	—	—	—
	CeO ₂	0.2	—	—	—	5.0	—	—	—	—
	Pr ₂ O ₃	0.2	—	—	—	—	5.0	—	—	—
	Nd ₂ O ₃	0.2	—	—	—	—	—	5.0	—	—
	Sm ₂ O ₃	0.2	—	—	—	—	—	—	5.0	—
	Eu ₂ O ₃	0.2	—	—	—	—	—	—	—	5.0
	Gd ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Tb ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Dy ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Ho ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Er ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Tm ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Yb ₂ O ₃	0.2	—	—	—	—	—	—	—	—
	Lu ₂ O ₃	0.2	—	—	—	—	—	—	—	—
Stability		○	○	○	○	○	○	○	○	○
Softening point (°C)		429	436	438	439	437	436	438	437	437
Thermal expansion coefficient (×10 ⁻⁷ /°C)		110	108	108	108	109	107	109	107	108
Water resistance(%)		0.2	0.2	0.1	0.1	0.1	0.2	0.1	0.2	0.2
Mechanical strength (MPa)		92	88	89	90	89	91	88	89	88

Table 34

		Example							
No.		293	294	295	296	297	298	299	300
Composition	SiO ₂	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1
(wt%)	B ₂ O ₃	6.6	6.6	6.6	6.6	6.6	6.6	6.6	6.6
	ZnO	10.3	10.3	10.3	10.3	10.3	10.3	10.3	10.3
	Bi ₂ O ₃	75.1	75.1	75.1	75.1	75.1	75.1	75.1	75.1
	Al ₂ O ₃	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Na ₂ O	1.7	1.7	1.7	1.7	1.7	1.7	1.7	1.7
	Gd ₂ O ₃	5.0	—	—	—	—	—	—	—
	Tb ₂ O ₃	—	5.0	—	—	—	—	—	—
	Dy ₂ O ₃	—	—	5.0	—	—	—	—	—
	Ho ₂ O ₃	—	—	—	5.0	—	—	—	—
	Er ₂ O ₃	—	—	—	—	5.0	—	—	—
	Tm ₂ O ₃	—	—	—	—	—	5.0	—	—
	Yb ₂ O ₃	—	—	—	—	—	—	5.0	—
	Lu ₂ O ₃	—	—	—	—	—	—	—	5.0
Stability		○	○	○	○	○	○	○	○
Softening point (°C)		437	436	438	437	436	438	437	437
Thermal expansion coefficient (×10 ⁻⁷ /°C)		107	109	108	108	108	109	107	108
Water resistance (%)		0.1	0.2	0.1	0.2	0.2	0.2	0.1	0.2
Mechanical strength (MPa)		88	89	89	91	89	90	91	92

Table 36

		Comparative Example							
No.		310	311	312	313	314	315	316	317
Composition	SiO ₂	1.1	1.1	1.1	1.1	1.1	1.1	1.1	1.1
(wt%)	B ₂ O ₃	6.6	6.6	6.6	6.6	6.6	6.6	6.6	6.6
	ZnO	10.1	10.1	10.1	10.1	10.1	10.1	10.1	10.1
	Bi ₂ O ₃	73.2	73.2	73.2	73.2	73.2	73.2	73.2	73.2
	Al ₂ O ₃	0.2	0.2	0.2	0.2	0.2	0.2	0.2	0.2
	Na ₂ O	1.8	1.8	1.8	1.8	1.8	1.8	1.8	1.8
	Gd ₂ O ₃	7.0	—	—	—	—	—	—	—
	Tb ₂ O ₃	—	7.0	—	—	—	—	—	—
	Dy ₂ O ₃	—	—	7.0	—	—	—	—	—
	Ho ₂ O ₃	—	—	—	7.0	—	—	—	—
	Er ₂ O ₃	—	—	—	—	7.0	—	—	—
	Tm ₂ O ₃	—	—	—	—	—	7.0	—	—
	Yb ₂ O ₃	—	—	—	—	—	—	7.0	—
	Lu ₂ O ₃	—	—	—	—	—	—	—	7.0
Stability		×	×	×	×	×	×	×	×
Softening point (°C)		—	—	—	—	—	—	—	—
Thermal expansion Coefficient (×10 ⁻⁷ /°C)		—	—	—	—	—	—	—	—
Water resistance (%)		—	—	—	—	—	—	—	—
Mechanical strength (MPa)		—	—	—	—	—	—	—	—

Sc₂O₃, Y₂O₃, La₂O₃, CeO₂, Pr₂O₃, Nd₂O₃, Sm₂O₃, Eu₂O₃, Gd₂O₃, Tb₂O₃, Dy₂O₃, Ho₂O₃, Er₂O₃, Tm₂O₃, Yb₂O₃ and Lu₂O₃ function to improve the mechanical strength of the glass composition.

As clearly shown in Tables 31 to 36, if the sealing member requires higher strength, it is necessary to use a glass composition containing 0.1 wt% or more of at least one of them. However, if the total content of them was 7.0 wt% or more, the glass composition was devitrified in the manufacture thereof, thereby the glass composition was not obtained with

stability. Accordingly, the total content of them is preferably 0.1 to 5 wt%.

Table 37 shows the glass compositions manufactured to make a study of the weight ratio of ZnO to B₂O₃ (ZnO/B₂O₃). The glass compositions of Nos. 319 to 321 are Examples and those of Nos. 318 and 322 are Comparative Examples.

Table 37

		Comparative Example	Example			Comparative Example
No.		318	319	320	321	322
Composition (wt%)	SiO ₂	2.2	2.3	1.1	1.2	1.4
	B ₂ O ₃	7.5	9.0	7.0	7.0	5.7
	ZnO	5.8	9.9	13.1	17.5	17.4
	Bi ₂ O ₃	83.4	75.8	76.3	72.9	72.9
	Al ₂ O ₃	—	0.5	0.1	0.2	0.6
	Na ₂ O	1.1	2.0	1.0	1.2	2.0
	CaO	—	—	0.5	—	—
	SrO	—	—	0.8	—	—
	La ₂ O ₃	—	—	0.1	—	—
	CeO ₂	—	0.5	—	—	—
ZnO/B ₂ O ₃		0.77	1.1	1.87	2.5	3.05
Stability		△	○	○	○	△
Softening point (°C)		423	430	427	430	419
Thermal expansion coefficient (×10 ⁻⁷ /°C)		120	113	108	102	109
Water resistance (%)		1.6	0.2	0.3	0.2	0.3
Mechanical strength (MPa)		78	79	80	78	76

As clearly shown in Table 37, when the weight ratio of ZnO to B₂O₃ is 0.77 or 3.05, the crystal precipitation is apt to occur in the glass in the heat treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity.

Accordingly, to obtain a stable glass composition that hardly causes the crystal precipitation, the weight ratio of ZnO to B_2O_3 is preferably controlled to 1.1 to 2.5.

Table 38 show the glass compositions manufactured to make a study of the weight ratio of Al_2O_3 to SiO_2 (Al_2O_3/SiO_2). The glass compositions of Nos. 323 to 325 are Examples and those of Nos. 326 and 327 are Comparative Examples.

Table 38

		Example			Comparative Example	
No.		323	324	325	326	327
Composition (wt%)	SiO_2	1.6	1.1	2.0	2.4	1.0
	B_2O_3	5.8	5.0	6.8	6.5	6.6
	ZnO	11.8	11.2	9.6	9.5	9.5
	Bi_2O_3	78.7	81.3	78.3	78.9	79.9
	Al_2O_3	—	0.2	1.0	1.8	2.0
	Li_2O	—	—	2.0	—	—
	Na_2O	—	—	—	0.9	1.0
	CaO	1.8	0.4	—	—	—
	SrO	—	0.7	—	—	—
	La_2O_3	—	0.1	0.3	—	—
	Gd_2O_3	0.3	—	—	—	—
Al_2O_3/SiO_2		0.0	0.18	0.5	0.75	2.0
Stability		○	○	○	△	△
Softening point (°C)		445	422	416	438	440
Thermal expansion Coefficient ($\times 10^{-7}/^{\circ}C$)		102	104	115	106	107
Water resistance (%)		0.1	0.2	0.2	0.2	0.2
Mechanical strength (MPa)		88	80	76	81	85

As clearly shown in Table 38, when the weight ratio of Al_2O_3 to SiO_2 is 0.75 or more, the crystal precipitation is apt to occur in the glass in the heat

treatment of the glass composition after the manufacture thereof, which may possibly cause a problem in maintaining hermeticity. Accordingly, to obtain a stable glass composition that hardly causes the crystal precipitation, it is important to limit the weight ratio of Al_2O_3 to SiO_2 , which is preferably controlled to 0.5 or less.

Table 39 shows lead-containing glass compositions manufactured as Comparative Examples 328 to 331 to make a performance comparison with the bismuth glass composition.

Table 39

		Comparative Example			
No.		328	329	330	331
Composition (wt%)	SiO_2	8.4	2.0	4.2	2.4
	B_2O_3	7.9	10.7	15.0	7.6
	ZnO	1.9	0.5	10.3	8.9
	PbO	80.3	85.2	70.5	78.6
	Al_2O_3	1.5	1.1	—	0.4
	K ₂ O	—	—	—	1.0
	CaO	—	—	—	1.1
	BaO	—	0.5	—	—
Stability		○	○	○	○
Softening point (°C)		408	396	425	378
Thermal expansion Coefficient ($\times 10^{-7}/^\circ\text{C}$)		96	100	91	103
Water resistance (%)		0.5	0.7	0.3	1.1
Mechanical strength (MPa)		85	76	84	72

As clearly shown in Table 39, the glass composition for forming the sealing member according to the present invention showed the softening point of 450°C or lower even if lead was not contained therein. Further, it also showed

mechanical strength equal to that of the lead-containing glass composition and was excellent in water resistance.

By adjusting the composition ratio in the glass composition of the present invention, a desired sealing member for a PDP was obtained.

Example 6

With the sealing member 16 comprising the glass composition manufactured in Example 5, a front plate 1 and a rear plate 8 were sealed, or an air hole 17 and a glass tube 18 were sealed to form a PDP. Except the sealing step, the manufacturing method was the same as that explained in the present embodiment.

Powder of the glass composition and a low-expansion ceramic filler were mixed in a predetermined weight ratio, to which nitrocellulose and isoamyl acetate were added to prepare a paste. The thus obtained paste containing the glass composition was applied to a sealing part 15 of the rear plate 8 by injection and provisionally baked at 350°C. Then, the front plate 1 and the rear plate 8 were stacked. A glass tube 18 was arranged at the position of an air hole 17 formed in the rear plate 8, to which the paste containing the glass composition was applied by injection. This was baked at a predetermined baking temperature for 30 minutes to perform the sealing. The baking temperature was set higher than the softening point of each glass composition by 50°C.

Table 40 shows the sealing glasses (Nos. A, B, C, D) as Examples of the present invention and a sealing member as Comparative Example (No. E), together with the composition ratio between the glass composition and the low-expansion ceramic filler comprising cordierite or zircon used in each sealing member, the thermal expansion coefficient and the baking temperature of the sealing member.

Between the sealing member 16 and the glass panel of the front plate 1 or the rear plate 8, at which the sealing is attempted, stress warpage occurs due to a difference in thermal expansion after the sealing. In order to reduce the stress warpage so as not to cause damages such as break, the thermal expansion coefficient of the sealing member 16 is preferably $60 \times 10^{-7}/^{\circ}\text{C}$ to $80 \times 10^{-7}/^{\circ}\text{C}$.

Table 40

	Example				Comparative Example
No.	A	B	C	D	E
Glass composition No.	203	220	244	324	329
Composition ratio in the sealing member (wt%)	—	—	—	—	—
Glass composition	60	60	55	50	55
Cordierite	40	—	45	—	45
Zircon	—	40	—	50	—
Thermal expansion coefficient ($\times 10^{-7}/^{\circ}\text{C}$)	72	71	68	69	68
Baking temperature ($^{\circ}\text{C}$)	460	491	480	472	446

After the sealing, any of the sealing members 16

of Examples allowed satisfactory bonding of the front plate 1 and the rear plate 8 in the same manner as those of Comparative Examples, and showed required hermeticity. As a result, a PDP capable of lighting was manufactured.

The low softening point glass compositions described in Example 5 have the thermal expansion coefficient far greater than these. Therefore, a sealing member having a desired thermal expansion coefficient can be obtained by optionally mixing a low-expansion ceramic filler therein. For that purpose, it is necessary to mix them in a weight ratio (low-expansion ceramic filler/glass composition) of 0.01 or more. However, if the weight ratio (low-expansion ceramic filler/glass composition) exceeds 4.0, the amount of the glass composition in the sealing member becomes too small, which makes the bonding itself impossible and the hermeticity and the strength deteriorate significantly. In view of this, the weight ratio of the low-expansion ceramic filler to the glass composition is preferably 0.01 to 4.0.

As the low-expansion ceramic filler, for example, may be used at least one selected from the group consisting of cordierite, willemite, forsterite, anorthite, zircon, mullite, β -eucryptite, β -spodumene, cristobalite, barium titanate, titanium oxide, tin oxide, aluminum oxide, zirconium oxide, quartz glass and refractory glass.

Industrial Applicability

According to the present invention, a bismuth glass composition is provided with an appropriate thermal expansion coefficient and mechanical strength, excellent water resistance and low working temperature. Since the bismuth glass composition does not contain lead, it causes less load to the environment.